

# Part B-1

Answer all questions in this part.

Directions (31–50): For each statement or question, record on your separate answer sheet the number of the word or expression that, of those given, best completes the statement or answers the question. Some questions may require the use of the 2011 Edition Reference Tables for Physical Setting/Chemistry.

31 What is the number of electrons in an  $\text{Al}^{3+}$  ion?

- (1) 10 (3) 3  
(2) 13 (4) 16

Al +13p

32 The valence electron of which atom in the ground state has the greatest amount of energy?

- (1) cesium (3) rubidium  
(2) lithium (4) sodium

33 The numbers of protons and neutrons in each of four different atoms are shown in the table below.

Protons and Neutrons in Four Different Atoms

Atom	Number of Protons	Number of Neutrons
A	8	8
D	9	9
E	9	10
G	10	10

Which two atoms represent isotopes of the same element?

- (1) A and D (3) E and D  
(2) A and G (4) E and G

34 Which elements have the most similar chemical properties?

- (1) boron and carbon  
(2) oxygen and sulfur  
(3) aluminum and bromine  
(4) argon and silicon

35 Which element reacts with oxygen to form ionic bonds?

- (1) calcium (3) chlorine  
(2) hydrogen (4) nitrogen

36 The table below gives the atomic mass and the abundance of the two naturally occurring isotopes of chlorine.

Naturally Occurring Isotopes of Chlorine

Isotopes	Atomic Mass of the Isotope (u)	Natural Abundance (%)
$^{35}\text{Cl}$	34.97	75.76
$^{37}\text{Cl}$	36.97	24.24

Which numerical setup can be used to calculate the atomic mass of the element chlorine?

- (1)  $(34.97 \text{ u})(75.76) + (36.97 \text{ u})(24.24)$   
(2)  $(34.97 \text{ u})(0.2424) + (36.97 \text{ u})(0.7576)$   
(3)  $(34.97 \text{ u})(0.7576) + (36.97 \text{ u})(0.2424)$   
(4)  $(34.97 \text{ u})(24.24) + (36.97 \text{ u})(75.76)$

37 Which general trends in first ionization energy and electronegativity values are demonstrated by Group 15 elements as they are considered in order from top to bottom?

- (1) The first ionization energy decreases and the electronegativity decreases.  
(2) The first ionization energy increases and the electronegativity increases.  
(3) The first ionization energy decreases and the electronegativity increases.  
(4) The first ionization energy increases and the electronegativity decreases.

38 An aluminum sample has a mass of 80.01 g and a density of 2.70 g/cm<sup>3</sup>. According to the data, to what number of significant figures should the calculated volume of the aluminum sample be expressed?

- (1) 1 (3) 3  
(2) 2 (4) 4

$$D = \frac{M}{V} ; V = \frac{M}{D}$$

$$\frac{80.01 \text{ (4)}}{2.70 \text{ (3)}}$$

# June 19 2013 Geometry Regents Answers

**A Loxley**



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### **Table of Contents June 19 2013 Geometry Regents Answers**

1. Understanding the eBook June 19 2013 Geometry Regents Answers
  - The Rise of Digital Reading June 19 2013 Geometry Regents Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying June 19 2013 Geometry Regents Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an June 19 2013 Geometry Regents Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from June 19 2013 Geometry Regents Answers
  - Personalized Recommendations
  - June 19 2013 Geometry Regents Answers User Reviews and Ratings
  - June 19 2013 Geometry Regents Answers and Bestseller Lists

5. Accessing June 19 2013 Geometry Regents Answers Free and Paid eBooks
  - June 19 2013 Geometry Regents Answers Public Domain eBooks
  - June 19 2013 Geometry Regents Answers eBook Subscription Services
  - June 19 2013 Geometry Regents Answers Budget-Friendly Options
6. Navigating June 19 2013 Geometry Regents Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - June 19 2013 Geometry Regents Answers Compatibility with Devices
  - June 19 2013 Geometry Regents Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of June 19 2013 Geometry Regents Answers
  - Highlighting and Note-Taking June 19 2013 Geometry Regents Answers
  - Interactive Elements June 19 2013 Geometry Regents Answers
8. Staying Engaged with June 19 2013 Geometry Regents Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers June 19 2013 Geometry Regents Answers
9. Balancing eBooks and Physical Books June 19 2013 Geometry Regents Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection June 19 2013 Geometry Regents Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine June 19 2013 Geometry Regents Answers
  - Setting Reading Goals June 19 2013 Geometry Regents Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of June 19 2013 Geometry Regents Answers
  - Fact-Checking eBook Content of June 19 2013 Geometry Regents Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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