

3. A  $0.12 \text{ mol dm}^{-3}$  sodium methanoate solution dissociates completely into its ions.

- (a) Formulate the equation, including state symbols, for the equilibrium reaction of the hydrolysis of the methanoate anion. [1]

*Sections 1 and 21 of the data booklet may be used for parts (b) to (e).*

- (b) Calculate the value of  $K_a$ , the acid dissociation constant at 298 K, for an aqueous solution of methanoic acid. [1]

- (c) Calculate the value of  $K_b$ , the base dissociation constant, for the conjugate base. [1]

*(This question continues on the following page)*

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