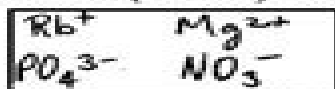


19. An aqueous solution of rubidium phosphate ( $\text{Rb}_3\text{PO}_4$ ) is mixed with an aqueous solution of magnesium nitrate and a precipitate forms.

a) Draw an (ion-box) with the four ions involved with this reaction. (1 mark)



b) Write a *balanced formula equation* for this reaction. Include all subscripts and coefficients. (2 marks)



c) Write a balanced, *complete (total) ionic equation* for this reaction. Include all subscripts, coefficients and ion charges where necessary. (2 marks)



d) Write a *net ionic equation* for this reaction. Include all subscripts, coefficients and ion charges where necessary. (2 marks)



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# Labpaq Answer Key Chemistry

**K Payea**



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