

Experiment 8 Report Sheet

Limiting Reactant

Date _____ Lab Sec. _____ Name _____ Desk No. _____

A. Precipitation of $\text{CaC}_2\text{O}_4 \cdot \text{H}_2\text{O}$ from the Salt Mixture

Unknown number _____

	Trial 1	Trial 2
1. Mass of beaker (g)	73.339 g	70.4301 g
2. Mass of beaker and salt mixture (g)	75.1309 g	71.4329 g
3. Mass of salt mixture (g)	1.5762 g	1.0028 g
4. Mass of filter paper (g)	0.3960 g	0.3962 g
5. Mass of filter paper and product after air-dried or oven-dried (g)	0.7609 g	0.70589 g
6. Mass of dried product (g)	0.3647 g	0.30969 g
7. Formula of dried product	CaC_2O_4	CaC_2O_4

B. Determination of Limiting Reactant

1. Limiting reactant in salt mixture (write complete formula)



2. Excess reactant in salt mixture (write complete formula)



Data Analysis

1. Moles of $\text{CaC}_2\text{O}_4 \cdot \text{H}_2\text{O}$ (or CaC_2O_4) precipitated (mol)	$2.496 \times 10^{-3} \text{ mol}$	$2.120 \times 10^{-3} \text{ mol}$
2. Moles of limiting reactant in salt mixture (mol)	1.980×10^{-3}	1.681×10^{-3}
• formula of limiting hydrate $\text{K}_2\text{C}_2\text{O}_4$		
3. Mass of limiting reactant in salt mixture (g)	0.3647 g	0.30969 g
4. Mass of excess reactant in salt mixture (g)	1.2118 g	0.6931 g
• formula of excess hydrate $\text{CaCl}_2 \cdot 2\text{H}_2\text{O}$		
5. Percent limiting reactant in salt mixture (%)	23.12 %	30.88 %
6. Percent excess reactant in salt mixture (%)	76.87 %	69.12 %
7. Mass of excess reactant that reacted (g)	0.2393 g	0.2456 g
8. Mass of excess reactant, unreacted (g)	0.9225 g	0.4475 g

*Show calculations for Trial 1 on next page.

Laboratory Manual Limiting Reactant

Daniel F McAuley



Laboratory Manual Limiting Reactant:

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