

CRIMINAL LAW 201 QUESTION PACK:

TOPIC – PERPETRATOR/ACCOMPLICE:

For a person to be guilty as an accomplice, there are certain requirements which must be complied with. Name and discuss these requirements. [7]

A person is an accomplice if

- 1) Although he does not comply with all the requirements for liability set out in the definition of the crime, and
- 2) Although the conduct required for a conviction is not imputed to him in terms of the doctrine of common purpose.

He engages in conduct whereby he furthers the commission of the crime by somebody else.

Requirements for liability as an accomplice

- 1) Act
- 2) Unlawfulness
- 3) Intention
- 4) Accessory character of liability

Explain the difference between a perpetrator, an accomplice and an accessory after the fact. In your answer you must also give an example of each. [8]

A person is a perpetrator if

(1) his conduct, the circumstances in which it takes place (including, where relevant, a particular description with which he as a person must, according to the definition of the offence, comply) and the culpability with which it is carried out are such that he satisfies all the requirements for liability contained in the definition of the offence

OR

(2) if, although his own conduct does not comply with that required in the definition of the crime, he acted together with one or more persons and the conduct required for a conviction is imputed to him by virtue of the principles relating to the doctrine of common purpose.

A person is an accomplice if:

(1) although he does not comply with all the requirements for liability set out in the definition of the crime, and

(2) although the conduct required for a conviction is not imputed to him in terms of the doctrine of common purpose, he engages in conduct whereby he furthers the commission of the crime by somebody else.

Matric Exam Paper Sa Criminal Law 201

Kathleen Armour



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 $2 \text{ CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) \rightarrow \text{C}_2\text{H}_2(\text{g}) + \frac{5}{2} \text{O}_2(\text{g}) \quad \Delta h = 94.5 \text{ kJ}$
 $\text{H}_2\text{O}(\text{g}) \rightarrow \text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \quad \Delta h = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{g}) \quad \Delta h = 283 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g})$

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 $\text{C}_2\text{H}_4(\text{g}) + 3 \text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + 2 \text{H}_2\text{O}(\text{l}) \quad \Delta h = 1411 \text{ kJ/mol}$
 $\text{C}_2\text{H}_6(\text{g}) + 7 \text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{l}) \quad \Delta h = 1560 \text{ kJ/mol}$
 $\text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l}) \quad \Delta h = 285 \text{ kJ/mol}$

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web thus for reaction 9.2 in which 1 mole of H_3O^+ and 1 mole of OH^- combine to form 2 moles of H_2O $\Delta H = -55.8 \text{ kJ}$ as shown note that the enthalpy change $\Delta H_{\text{products}} - \Delta H_{\text{reactants}}$ is positive if heat is absorbed that is if $\Delta H_{\text{products}}$

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