

JEE 2023 Session-1 24th Jan to 1st Feb 2023

Application No	
Candidate Name	
Roll No	
Test Date	29/01/2023
Test Time	9:00 AM - 12:00 PM
Subject	B TECH

Section : Physics Section A

Q.1 Match List I with List II:

List I (Physical Quantity)	List II (Dimensional Formula)
A. Pressure gradient	I. $[M^2L^2T^{-2}]$
B. Energy density	II. $[M^2L^{-1}T^{-2}]$
C. Electric Field	III. $[M^2L^{-2}T^{-2}]$
D. Latent heat	IV. $[M^2L^2T^{-2}A^{-1}]$

Choose the *correct* answer from the options given below:

- Options:
1. A-II, B-III, C-IV, D-I
 2. A-III, B-II, C-IV, D-I
 3. A-II, B-III, C-I, D-IV
 4. A-III, B-II, C-I, D-IV

Question Type : MCQ
 Question ID : 3666942007
 Option 1 ID : 3666946202
 Option 2 ID : 3666946204
 Option 3 ID : 3666946201
 Option 4 ID : 3666946203
 Status : Answered
 Chosen Option : 2

Q.2 A car is moving on a horizontal curved road with radius 50 m. The approximate maximum speed of car will be, if friction between tyres and road is 0.34. [take $g = 10 \text{ ms}^{-2}$]

- Options:
1. 22.4 ms^{-1}
 2. 13 ms^{-1}
 3. 3.4 ms^{-1}
 4. 17 ms^{-1}

Question Type : MCQ
 Question ID : 3666942017
 Option 1 ID : 3666946244
 Option 2 ID : 3666946243
 Option 3 ID : 3666946241
 Option 4 ID : 3666946242
 Status : Answered
 Chosen Option : 3

Jee Paper 2 Code K 2014 Answer Pics

Ian Pickup



Jee Paper 2 Code K 2014 Answer Pics:

The Enigmatic Realm of **Jee Paper 2 Code K 2014 Answer Pics**: Unleashing the Language is Inner Magic

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the objectives of the boiling point determination lab were to determine the boiling point of a mystery substance provided to each student boiling point is the point at which a liquid changes physical states and becomes a vapor or gas boiling point is

important to understand because it can be used to help distinguish between different substances

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the boiling point of an organic liquid is the temperature at which its vapour pressure equals the atmospheric pressure over the liquid or it is the temperature at which the vapour and liquid phases are in equilibrium at a given pressure the boiling point is considered as a criterion of purity of a compound and is useful for identification of

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given that the boiling points of ethanol and methanol 78 °C and 64 °C respectively are far off in a significant way this seems like a valid conclusion conclusion thermometer calibrations are necessary in order to collect reliable information as precise calculations are imperative to gathering correct results

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melting point 1 a push the open end of a capillary tube into the powdered acetamide b move the powder to the closed end of the capillary tube by tapping it on the table repeat until the powdered acetamide occupies 1–2 mm of the capillary tube end c

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discussion and conclusion the technique recrystallization was used because it is a simple method to purify a solid compound the percentage yield for benzoic acid and acetanilide acid being high indicated the initial sample had a low concentration of impurities with the melting point being in a specific

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determination of boiling points introduction the boiling point of a compound is the temperature at which it changes from a liquid to a gas this is a physical property often used to identify substances or to check the purity of the compound it

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concluded that the unknown substance 193 is isopropyl alcohol on account of its observed boiling point of 88 c

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conclusion after completing this experiment the melting point of three different unknown compounds was obtained using the mel temp apparatus compound a had a melting point range of 114 6 115 5 c thus revealing it to be acetanilide compound b had a melting point range of 122 6 124 5 c thus

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melting point can be used as the identity and purity of an organic solid however it is not the same as the boiling point the boiling point is a characteristic unique of pure liquids the boiling point is the temperature at which the vapor pressure of a liquid is equal to the atmospheric pressure

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aug 11 2020 boiling points are very sensitive to changes in applied pressure so all boiling points should be reported with the measured pressure a compound s normal boiling point refers to its boiling point at a pressure of 760 text mm ce hg

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the purpose of the experiment is to understand what vapour pressure is and how it changes with temperature to understand the relationship between vapour pressure at a given temperature and boiling temperature at a given pressure to determine the vapour pressure of a pure liquid at various temperatures

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exploring boiling points abstract in this experiment we are given a set of straight chain alkanes and a set of compounds to compare by following the trendline of the straight chain alkanes it could be concluded that as the molar mass of the straight chain alkanes increases their boiling points increase as well

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laboratory partners samantha and mackenzie goals □ correctly determine the boiling point of an unknown liquid □ measure the refractive index of an unknown □ use the gas chromatogram to determine the mass percentage of a mixture and

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50 c so in order to convert an experimental measurement taken in calgary which is at higher altitude than sea level and so is at a lower pressure to that reported for sea level higher pressure one needs to add a correction factor since the boiling point at sea level is higher than that at higher altitudes

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there are a variety of methods by which a sample s boiling point can be determined including distillation reflux and by using a thiele tube the most straightforward method uses a thiele tube and has the advantage of using less than 0 5ml 0 5 ml of material

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