

2.3. $\Delta H = \Delta H_{\text{f}}^{\circ}(\text{C}_6\text{H}_6) - 6 \Delta H_{\text{f}}^{\circ}(\text{C})$

When you have the needed data to calculate the enthalpy of formation of benzene, you can calculate the enthalpy of formation of benzene.

The same is valid for the formation of the all the gas fuels, but make it clear that you have the preferred along with fuel.

Why do we have to calculate the enthalpy of formation?



1. An 80-year-old male.

Calculate the $\Delta H_{\text{f}}^{\circ}$ of benzene.

Types for the fuel: benzene.

Standard $\Delta H_{\text{f}}^{\circ}$ of C_6H_6 is

100 kcal/mol.

100 kcal/mol.

2. Calculate the $\Delta H_{\text{f}}^{\circ}$ of benzene.



The chemistry from the cell they the point not to calculate up to the point.

Point 1: 100 kcal/mol

Point 2: 100 kcal/mol

Point 3: 100 kcal/mol

3. Calculate the $\Delta H_{\text{f}}^{\circ}$ of benzene.

1. $\Delta H_{\text{f}}^{\circ}$ of benzene, 100 kcal/mol.

2. $\Delta H_{\text{f}}^{\circ}$ of benzene, 100 kcal/mol.

3. $\Delta H_{\text{f}}^{\circ}$ of benzene, 100 kcal/mol.

4. $\Delta H_{\text{f}}^{\circ}$ of benzene, 100 kcal/mol.

5. $\Delta H_{\text{f}}^{\circ}$ of benzene, 100 kcal/mol.

6. $\Delta H_{\text{f}}^{\circ}$ of benzene, 100 kcal/mol.

7. $\Delta H_{\text{f}}^{\circ}$ of benzene, 100 kcal/mol.

8. $\Delta H_{\text{f}}^{\circ}$ of benzene, 100 kcal/mol.

9. $\Delta H_{\text{f}}^{\circ}$ of benzene, 100 kcal/mol.

10. $\Delta H_{\text{f}}^{\circ}$ of benzene, 100 kcal/mol.

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