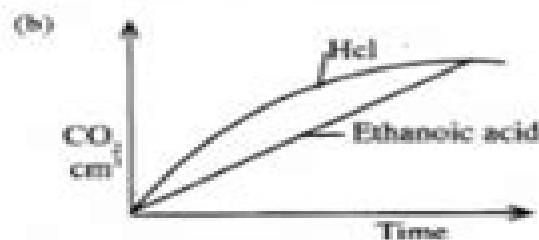


(b) Moles of brown gas (NO_2) = $\frac{0.29}{24} = 0.012 \checkmark$

Moles of lead (II) nitrate = $\frac{1}{2} \times \frac{0.29}{24} = 0.006 \checkmark$ (2 marks)

Mass of lead (II) nitrate = $0.006 \times 331 \checkmark$
 = $1.9998 \text{ g} \checkmark$

8. (a) Strong acid ionises fully. (1 mark)



(2 marks)

9. (a) Hydrogen is expensive.
 Hydrogen is explosive. (2 marks)

10. (a) • Green colour of chlorine disappears.
 • Brown gas is produced or black solid is deposited. (1 mark)



Explanation: Iodine oxidation state changes from -1 to 0 hence oxidation while chlorine oxidation state changes from 0 to -1 hence reduction. (1 mark)

11. (a) Carbon (II) oxide is formed in the internal combustion engines when fuel burns under limited oxygen. (1 mark)

(b) Pollutant gas - Carbon (IV) oxide, Nitrogen (IV) oxide and Sulphur (IV) oxide.
 (Any two) (2 marks)

12. (a) • Small piece of sodium metal (pea size) with alot of water.
 • Perform the experiment wearing goggles. (1 mark)

(b) Electrolysis. (1 mark)

(c) Manufacture of soap. (1 mark)

13. Deliquescent substance absorbs water from the atmosphere to form a solution, while a fluorescent substance loses water of crystallisation to the atmosphere. (2 marks)

14. P is in alkanol R - OH. The alkanol reacts with sodium metal to produce the colourless gas.

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