

Name _____ Date _____ Class _____

CHAPTER 12

STUDY GUIDE FOR CONTENT MASTERY

Section 12.3 Limiting Reactants

In your textbook, read about why reactions stop and how to determine the limiting reactant.

Study the diagram showing a chemical reaction and the chemical equation that represents the reaction. Then complete the table. Show your calculations for questions 15–17 in the space below the table.



The molar masses of O_2 , N_2 , and NO_2 are 32.00 g/mol, 28.02 g/mol, and 46.01 g/mol, respectively.

Amount of O_2	Amount of N_2	Amount of NO_2	Limiting Reactant	Amount and Mass of Excess Reactant
1 molecule	2 molecules	2 molecules	none	none
4 molecules	4 molecules	4 molecules	N_2	2 molecules O_2
2 molecules	4 molecules	4 molecules	O_2	2 molecules N_2
1.00 mol	2.00 mol	2.00 mol	none	none
4.00 mol	4.00 mol	4.00 mol	N_2	2.00 mol O_2
2.00 mol	4.00 mol	4.00 mol	O_2	2.00 mol N_2
1.00 mol	4.00 mol	4.00 mol	N_2	2.00 mol O_2
4.000 mol	2.000 mol	4.000 mol	N_2	2.000 mol O_2
32.00 g	56.04 g	92.02 g	none	none
64.00 g	56.04 g	92.02 g	O_2	32.00 g N_2
64.00 g	112.08 g	184.04 g	O_2	32.00 g N_2

balanced equation mole ratio = 2 mol NO_2 /1 mol O_2

16.00 g O_2 \times 1 mol O_2 /32.00 g O_2 = 0.500 mol O_2

28.02 g N_2 \times 1 mol N_2 /28.02 g N_2 = 1.000 mol N_2

actual mole ratio = 0.500 mol NO_2 /0.500 mol O_2 = 1.000 mol NO_2 /1.000 mol O_2

Because the actual mole ratio of NO_2 - O_2 is larger than the balanced equation mole ratio of NO_2 - O_2 , there is an excess of N_2 ; O_2 is the limiting reactant.

Mass of NO_2 used = 0.500 mol NO_2 \times 46.01 g/mol NO_2 = 23.01 g NO_2

0.500 mol NO_2 \times 30.01 g/mol NO_2 = 15.01 g NO_2

Mass of NO_2 produced = 0.500 mol NO_2 \times 46.01 g/mol NO_2 = 23.01 g NO_2

Excess N_2 = 28.02 g N_2 - 15.01 g N_2 = 13.01 g N_2

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CHAPTER 13

STUDY GUIDE FOR CONTENT MASTERY

Section 13.4 Percent Yield

In your textbook, read about the yield of products.

Study the diagram and the example problem.



Example Problem: The following chemical equation represents the production of palladium, a valuable metal in the manufacturing of some semiconductor devices.



In one experiment, the reaction yielded 7.42 g of the metal from a 7.00-g sample of palladium. Determine the percent yield of this reaction. The molar masses of Ga and Ga_2O_3 are 69.72 g/mol and 157.44 g/mol, respectively.

Use the information in the diagram and example problem to evaluate each value or expression below. If the value or expression is correct, write correct. If it is incorrect, write the correct value or expression.

- actual yield unknown 7.42 g Ga_2O_3
- mass of reactant: 7.00 g Ga correct
- number of moles of reactant: 7.00 g Ga \times $\frac{157.44 \text{ g } Ga_2O_3}{1 \text{ mol } Ga}$ 7.00 g Ga \times 1 mol Ga /69.72 g Ga
- number of moles of product: 0.100 mol Ga \times $\frac{2 \text{ mol } Ga_2O_3}{1 \text{ mol } Ga}$ 0.100 mol Ga \times 2 mol Ga_2O_3 /1 mol Ga
- theoretical yield: 0.100 mol Ga_2O_3 \times $\frac{157.44 \text{ g } Ga_2O_3}{1 \text{ mol } Ga_2O_3}$ correct
- percent yield: $\frac{7.42 \text{ g } Ga_2O_3}{7.42 \text{ g } Ga_2O_3} \times 100$ 7.42 g Ga_2O_3 /7.42 g Ga_2O_3 \times 100

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Table of Contents Mcgraw Hill Ch 14 Answer Key Chemistry

1. Understanding the eBook Mcgraw Hill Ch 14 Answer Key Chemistry
 - The Rise of Digital Reading Mcgraw Hill Ch 14 Answer Key Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Mcgraw Hill Ch 14 Answer Key Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mcgraw Hill Ch 14 Answer Key Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mcgraw Hill Ch 14 Answer Key Chemistry
 - Personalized Recommendations
 - Mcgraw Hill Ch 14 Answer Key Chemistry User Reviews and Ratings
 - Mcgraw Hill Ch 14 Answer Key Chemistry and Bestseller Lists

5. Accessing Mcgraw Hill Ch 14 Answer Key Chemistry Free and Paid eBooks
 - Mcgraw Hill Ch 14 Answer Key Chemistry Public Domain eBooks
 - Mcgraw Hill Ch 14 Answer Key Chemistry eBook Subscription Services
 - Mcgraw Hill Ch 14 Answer Key Chemistry Budget-Friendly Options
6. Navigating Mcgraw Hill Ch 14 Answer Key Chemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - Mcgraw Hill Ch 14 Answer Key Chemistry Compatibility with Devices
 - Mcgraw Hill Ch 14 Answer Key Chemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mcgraw Hill Ch 14 Answer Key Chemistry
 - Highlighting and Note-Taking Mcgraw Hill Ch 14 Answer Key Chemistry
 - Interactive Elements Mcgraw Hill Ch 14 Answer Key Chemistry
8. Staying Engaged with Mcgraw Hill Ch 14 Answer Key Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mcgraw Hill Ch 14 Answer Key Chemistry
9. Balancing eBooks and Physical Books Mcgraw Hill Ch 14 Answer Key Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mcgraw Hill Ch 14 Answer Key Chemistry
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Mcgraw Hill Ch 14 Answer Key Chemistry
 - Setting Reading Goals Mcgraw Hill Ch 14 Answer Key Chemistry
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Mcgraw Hill Ch 14 Answer Key Chemistry
 - Fact-Checking eBook Content of Mcgraw Hill Ch 14 Answer Key Chemistry
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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