

Chemistry – Unit 5 Worksheet 1

Relative Masses of the Elements

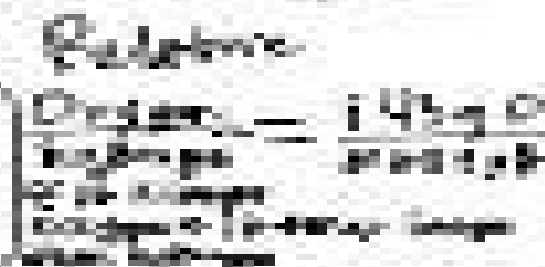
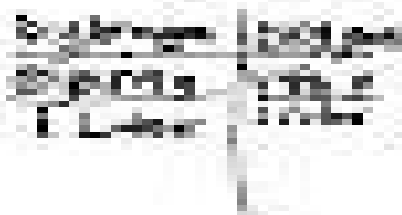
Relative Mass From Counts

We have established that atoms of hydrogen and oxygen are present in the compound in the same proportions as:

1. Equal volumes of gases contain the same number of molecules as the other gas (Avogadro's hypothesis) **Simplest Equation**
2. Atoms of oxygen are 16 times as heavy as those of hydrogen (relative mass).



Using the relative mass of oxygen, Avogadro's hypothesis explains how that the density of oxygen gas is equivalent to hydrogen and propane in 1:1.1 ratio, characteristic among other things gas could be distinguished in 1899 which. This shows that molecules of O_2 are 16 times as heavy as those molecules of hydrogen. **From H₂O ratio**



Avogadro's hypothesis that molecules were able to determine the relative masses of all the elements using this technique. Measurements of the density of the gases of all many other elements are difficult, it was impossible. However, it is now possible to determine relative masses of all the particles (atoms, molecules, ions) by using mass spectrometry.

Relative Mass From Composition

Many substances contain both oxygen and hydrogen (e.g. water). The relative masses of the particles are the same as those of the atoms (e.g. 16 for oxygen, 1 for hydrogen). This means that the composition of the substance can be used to determine the relative masses of the particles. For example, if a substance contains 8 g of oxygen and 1 g of hydrogen, then the relative mass of the particles is 16:1.

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