Chemistry Unit 8 Worksheet 4 Samples of Every Kind of Problem

On a separate sheet of paper, write a complete solution to each of the problems below. Follow the procedure outlined in class. Be sure to circle your final answer.

 Calculate the number of moles of potassium chlorate, KClO₃ (s), that must decompose to produce potassium chloride, KCl (s), and 1.8 moles of oxygen gas.

Equation:	2 KClO ₃ →	2 KCI +	3 O2
Before:	1.2	0	0
Change:	-1.2	+1.2	+1.8
After:	0	1.2	1.8

1.2 moles KClO₃ needed

2. In a single displacement reaction, magnesium metal reacts with hydrochloric acid to produce magnesium chloride and hydrogen gas. How many moles of hydrochloric acid are needed to completely react with 2.43 g of magnesium?

Equation:	Mig +	2 HCl ->	MgCl ₂ +	H ₂
Before:	0.100	- 7	0	0
Change:	-0.100	-0.200 +	0.160 +	0.100
Afters	0		0.100	0.100

0,200 moles HCI needed

0.73 g x 1 mole = .0040 mole 183 g

3. Ethane, C₂H₆ reacts with oxygen gas to produce carbon dioxide gas and water vapor. What mass of oxygen gas is required to react with 2.20 moles of ethane?

Equation:	2 C2H6	+ 7 O ₂	→ 4 CO₂	+ 6 H ₂ O
Before:	2.20	36.56	0	0
Change:	-2.20	- 7.70	+ 4,40	+ 6.60
After:	0	X 5	4.40	6.60

O₂ needed: 7.70 mole x 32.0 g = 246 g O₂ required 1 mole

 Determine the mass of sodium nitrate produced when 0.73 g of nickel (II) nitrate reacts with sodium hydroxide according to the following equation: Ni(NO3)2 + 2 NaOH --> Ni(OH)2 + 2 NaNO3

 Equation:
 Ni(NO₃)₂ + 2 NaOH
 → Ni(OH)₂ + 2 NaNO₃

 Before:
 0.0040
 XS
 0
 0

 Change:
 -0.0040
 -0.0080
 +0.0040
 +0.0080

 After:
 0
 XS
 0.0040
 0.0080

NaNO, produced: 0.0080 mole x 85.0 g = 0.68 g 1 mole

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