

Model 1: Rules of Solubility in Aqueous Solutions

Several solid compounds are placed in water to determine if they are soluble. The results are shown in the chart below. If it does dissolve no mark is made.

The top row shows the cation of the compound. The far left column shows the anion of the compound. For example, $\text{Mg}(\text{OH})_2$ is insoluble, MgBr_2 is soluble and AgCl is insoluble.

Table 1 – This table presents an overview of solubility of selected salts in water

X = does not dissolve in water. * partially soluble.

	NH_4^+	Li^+	Na^+	K^+	Mg^{2+}	Ca^{2+}	Sr^{2+}	Ba^{2+}	Ag^+	Pb^{2+}	Hg_2^{2+}	Fe^{3+}	Cu^{2+}	Zn^{2+}
NO_3^-														
$\text{C}_2\text{H}_3\text{O}_2^-$														
CO_3^{2-}														
CO_3^{2-}														
Cl^-									X	X	X			
Br^-									X	X	X			
I^-									X	X	X			
SO_4^{2-}						X	X	X	X	X	X			
OH^-					X	*	*		X	X	X	X	X	X
S^{2-}					X	X	X	X	X	X	X	X	X	X
CO_3^{2-}					X	X	X	X	X	X	X	X	X	X
PO_4^{3-}					X	X	X	X	X	X	X	X	X	X
$\text{C}_2\text{O}_4^{2-}$					X	X	X	X	X	X	X	X	X	X
CrO_4^{2-}					X	X	X	X	X	X	X	X	X	X

Key Questions

1. Is calcium carbonate soluble or insoluble?
2. Is silver bromide soluble or insoluble?
3. Is iron (III) sulfate soluble or insoluble?
4. For what cations are the compounds always soluble in water?
5. For what anions are the compounds always soluble in water?
6. For what anions are most of the compounds usually soluble (with some exceptions)?
7. For those anions that usually form soluble compounds, which cations result in the

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