

Chemistry Unit 5 - Empirical Formula Lab

Introduction

In this experiment, you will use a known amount of a solid hydrocarbon with hydrogens and 80% oxygen and the relative amounts of the two elements to determine the empirical formula of the hydrocarbon. The empirical formula of a compound is the simplest whole number ratio of the atoms of each element in the compound. For example, the empirical formula of glucose is $C_6H_{12}O_6$. The empirical formula of glucose is CH_2O .

Pre-Lab Questions:

1. Read through the following procedure carefully and make sure you understand the steps before you start the lab. Underline words that should appear in the procedure. Write down:
 - 1. Purpose of the experiment
 - 2. Materials and equipment
 - 3. Safety hazards
 - 4. Pre-lab questions
 2. Describe your experimental procedure in your own words. Be sure to include the following information:
 - a. How you will measure the mass of the hydrocarbon.
 - b. How you will measure the mass of the oxygen.
 - c. How you will determine the empirical formula.
- Write your answers to the pre-lab questions in your notebook on the lab.

Procedure - Part 1

1. Weigh a small amount (about 0.5 g) of the hydrocarbon in a pre-weighed crucible. Find the mass of the hydrocarbon and record it.
2. Add a small amount of oxygen (about 0.5 g) to the crucible.
3. Heat the crucible in a Bunsen burner until the hydrocarbon is completely oxidized.

Procedure - Part 2

1. Cool the crucible and weigh it. Calculate the mass of the oxygen that was added to the hydrocarbon.
2. Calculate the mass of the hydrocarbon from the mass of the hydrocarbon and the mass of the oxygen that was added to it. Use the mass of the hydrocarbon and the mass of the oxygen to determine the empirical formula of the hydrocarbon.
3. Repeat the experiment with a different amount of the hydrocarbon and oxygen. Compare your results to the results of the first experiment.
4. Repeat the experiment with a different amount of the hydrocarbon and oxygen. Compare your results to the results of the first experiment.

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