

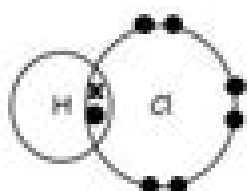
C3 - STRUCTURE AND BONDING 3.5 Covalent bond

Aiming for Grade 4

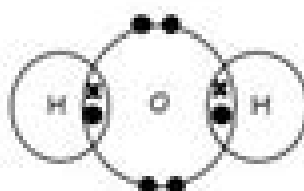
Ionic bonds form between **metals** and **non-metals**. Covalent bonds form between **non-metals**. Covalent bonds form when atoms **share** pairs of **electrons**. These bonds are very **strong**.

Share, Metals, strong, non-metals, electrons
You can use the words more than once.

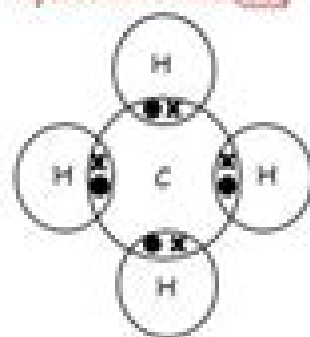
Name these small molecules and state their formulae.



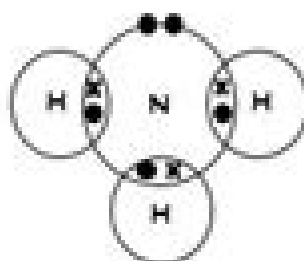
Hydrochloric Acid (HCl)



Water (H_2O)



Methane (CH_4)



Ammonia (NH_3)

Assessment

Aiming for Grade 6

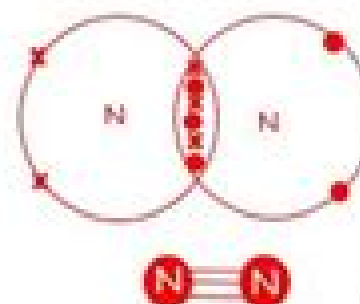
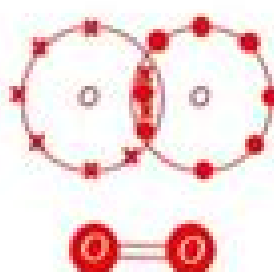
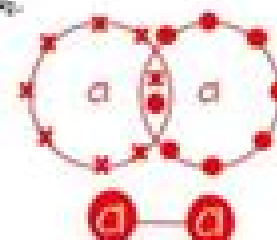
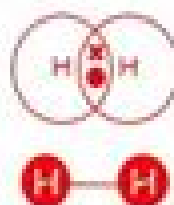
Describe how Hydrogen and Chlorine form a covalent bond to form Hydrochloric acid.

To gain full outer shells Hydrogen needs to gain an electron and Chlorine needs to lose one. Hydrogen shares its outer shell electron with Chlorine and Chlorine shares one of its electrons.

What is a double covalent bond?

Two shared pairs of electrons

Draw dot and cross diagrams and ball and stick diagrams for H_2 , Cl_2 , O_2 , and N_2 .



Assessment

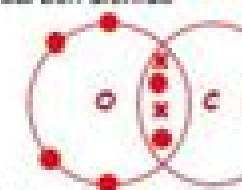
Aiming

Draw ball and stick or the following covalent

Three hydrogen atoms
one phosphorus (phos)



Carbon dioxide



Ethane (2 carbon with hydrogen)



How do double and triple properties?

A double or triple bond the bond making the me

Assessment
Grade 6

1. I can draw dot and

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Table of Contents June 2013 Aqa Chemistry C3 Mark Scheme

1. Understanding the eBook June 2013 Aqa Chemistry C3 Mark Scheme
 - The Rise of Digital Reading June 2013 Aqa Chemistry C3 Mark Scheme
 - Advantages of eBooks Over Traditional Books
2. Identifying June 2013 Aqa Chemistry C3 Mark Scheme
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an June 2013 Aqa Chemistry C3 Mark Scheme
 - User-Friendly Interface
4. Exploring eBook Recommendations from June 2013 Aqa Chemistry C3 Mark Scheme
 - Personalized Recommendations
 - June 2013 Aqa Chemistry C3 Mark Scheme User Reviews and Ratings
 - June 2013 Aqa Chemistry C3 Mark Scheme and Bestseller Lists
5. Accessing June 2013 Aqa Chemistry C3 Mark Scheme Free and Paid eBooks

- June 2013 Aqa Chemistry C3 Mark Scheme Public Domain eBooks
 - June 2013 Aqa Chemistry C3 Mark Scheme eBook Subscription Services
 - June 2013 Aqa Chemistry C3 Mark Scheme Budget-Friendly Options
6. Navigating June 2013 Aqa Chemistry C3 Mark Scheme eBook Formats
- ePub, PDF, MOBI, and More
 - June 2013 Aqa Chemistry C3 Mark Scheme Compatibility with Devices
 - June 2013 Aqa Chemistry C3 Mark Scheme Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of June 2013 Aqa Chemistry C3 Mark Scheme
 - Highlighting and Note-Taking June 2013 Aqa Chemistry C3 Mark Scheme
 - Interactive Elements June 2013 Aqa Chemistry C3 Mark Scheme
8. Staying Engaged with June 2013 Aqa Chemistry C3 Mark Scheme
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers June 2013 Aqa Chemistry C3 Mark Scheme
9. Balancing eBooks and Physical Books June 2013 Aqa Chemistry C3 Mark Scheme
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection June 2013 Aqa Chemistry C3 Mark Scheme
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine June 2013 Aqa Chemistry C3 Mark Scheme
- Setting Reading Goals June 2013 Aqa Chemistry C3 Mark Scheme
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of June 2013 Aqa Chemistry C3 Mark Scheme
- Fact-Checking eBook Content of June 2013 Aqa Chemistry C3 Mark Scheme
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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