(e) Compound A is an orange ionic compound of chromium with the percentage composition by mass N, 11.11%; H, 3.17%; Cr, 41.27%; O, 44.45%. Compound A does not have water of crystallisation.

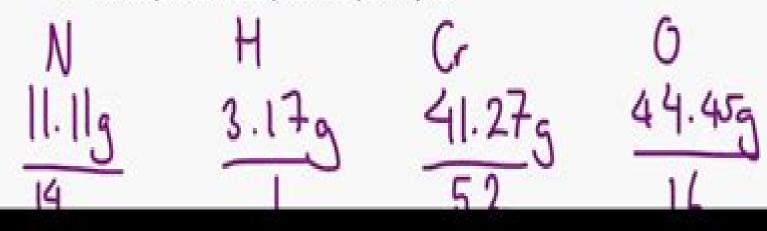
On gentle heating, compound A decomposes to form three products, B, C and water.

B is a green oxide of chromium with a molar mass of 152.0 gmol⁻¹.

C is a gas. ATHIP, each cook decimetre of C has a mass of 1.17g.

In the steps below, show all your working.

- Calculate the empirical formula of compound A.
- Deduce the ions that make up the ionic compound A.
- dentity substances B and C.
- Write an equation for the decomposition of compound A by heat.



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