




Molecular Models: Lewis Electron Dot Structures and Molecular Geometry

Part A: Use the molecular model kit and complete the following table.

Molecular formula	Total number of valence electrons	Lewis electron dot structure	Sketch of three-dimensional geometry	Name of the molecular geometry	Molecule is polar or nonpolar?
N_2	$5 + 5 = 10$	$:N \equiv N:$	$N \equiv N$	Linear	Non polar
NH_3	$5 + 3 \times 1 = 8$	$H \cdot \cdot \ddot{N} \cdot \cdot H$ $\quad $ $\quad H$		Trigonal Pyramidal	Polar
H_2S	$2 \times 1 + 6 = 8$	$H : \ddot{S} : H$ $\quad \cdot \cdot$		Bent	Polar
CO_2	$4 + 12 = 16$	$\ddot{O} = C = \ddot{O}$	$O = C = O$	linear	non Polar
PF_3	$5 + 21 = 26$	$\begin{array}{c} \cdot \cdot \\ \cdot \cdot \\ \cdot \cdot \end{array} \ddot{F} - \ddot{P} - \ddot{F} \cdot \cdot$ $\quad \cdot \cdot$ $\quad \cdot \cdot$		Trigonal Pyramidal	Polar

Lewis Structures And Molecular Geometry Guided Inquiry Kit Answers

Jeffrey R. Buell



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, 1968 **Exploring Osaka** David M. Dunfield, 1993 A comprehensive English language guide to the city of Osaka for business travellers and tourists alike Second only to Tokyo as a banking and trade centre of Japan Osaka is a centre of tradition and culture as well the home of bunraku puppet theatre for example and is minutes by rail from Japan's ancient capitals of Kyoto and Nara A variety of suggested sightseeing itineraries are included as well as recommended museums restaurants and hotels a list of festival dates and sources of additional information *Lewis Structures and Molecular Geometry* Jeffrey R. Buell, 1992 *Lewis Structures and Molecular Geometry (computer File)*. Jeffrey R. Buell, 1968

Molecules and Models Arne Haaland, 2008-03-06 This book provides a systematic description of the molecular structures and bonding in simple compounds of the main group elements with particular emphasis on bond distances bond energies and coordination geometries The description includes the structures of hydrogen halogen and methyl derivatives of the elements in each group some of these molecules are ionic some polar covalent The survey of molecules whose structures conform to well established trends is followed by representative examples of molecules that do not conform We also describe electron donor acceptor and hydrogen bonded complexes Chemists use models to systematize our knowledge to memorize information and to predict the structures of compounds that have not yet been studied The book provides a lucid discussion of a number of models such as the Lewis electron pair bond and the VSEPR models the spherical and polarizable ion models and molecular orbital calculations and it outlines the successes and failures of each

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