# IB SL and HL Topic 1 to 5 Chemistry MCQ 40 Marks

- What amount of oxygen, O<sub>2</sub>, (in moles) contains 1.8×10<sup>22</sup> molecules?
  - A. 0.0030
  - B. 0.030
  - C. 0.30
  - D. 3.0

2. Which compound has the empirical formula with the greatest mass?

- A. C<sub>2</sub>H<sub>6</sub>.
- B. C<sub>4</sub>H<sub>10</sub>
- C.  $C_5H_{10}$
- D. C<sub>6</sub>H<sub>6</sub>

(Total I mark)

(Total I mark)

3.  $C_2H_2(g) + O_2(g) \rightarrow CO_2(g) + H_2O(g)$ 

When the equation above is balanced, what is the coefficient for oxygen?

- Λ. 2
- B. 3
- C. 4
- D. 5

(Total I mark)

3.0 dm<sup>3</sup> of sulfur dioxide is reacted with 2.0 dm<sup>3</sup> of oxygen according to the equation below.

$$2SO_2(g) + O_2(g) \rightarrow 2SO_3(g)$$

What volume of sulfur trioxide (in dm<sup>3</sup>) is formed? (Assume the reaction goes to completion and all gases are measured at the same temperature and pressure.)

- A. 5.0
- B. 4.0
- C. 3.0
- D. 2.0

(Total I mark)

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