

## Chemistry Semester 1 Final Exam Review

- 1.
  2. What is a definition of Chemistry?
  3. Matter includes all of the following EXCEPT:
    - a. air
    - b. light
    - c. smoke
    - d. water vapor
  4. What is matter?
  5. A physical change occurs when a
    - a. peach spoils
    - b. bracelet turns your wrist green
    - c. copper bowl tarnishes
    - d. glue gun melts a glue stick
  6. The particles in a solid are
    - a. packed closely together
    - b. very far apart
    - c. constantly in motion
    - d. able to slide past one another
- Of the following groups, 1, 2, 14, 15, 17, 18:
7. Group(s) \_\_\_\_\_ contains only metals
  8. Group (s) \_\_\_\_\_ is called a Noble gases
  9. Group(s) \_\_\_\_\_ has 7 electrons on the valence shell
  10. Group(s) \_\_\_\_\_ needs to gain 3 electrons to fill a stable octet
  11. Group(s) \_\_\_\_\_ has a  $2+$  oxidation state
  12. Group(s) \_\_\_\_\_ has an electron dot notation of  $\bullet X \bullet$
  13. Using the periodic table, an element similar to carbon would be \_\_\_\_\_
  14. Based on their location on the periodic table, you could infer that \_\_\_\_\_ are very unreactive
  15. The horizontal row on the periodic table is called a(n) \_\_\_\_\_
  16. The vertical column of blocks on the periodic table is called a(n) \_\_\_\_\_
  17. The elements on the zigzag line in the periodic table are \_\_\_\_\_
  18. Which is NOT a property of metal?
    - a. malleability
    - b. ability to lose electrons
    - c. unreactivity
    - d. ability to conduct heat
  19. A volume of 1 cc ( $\text{cm}^3$ ) is the same as \_\_\_\_\_ mL
  20. To determine density, the quantities that must be measured are \_\_\_\_\_ and \_\_\_\_\_
  21. The relationship between the mass  $m$  of a material, its volume  $V$ , and its density  $D$  is
    - a.  $V = mD$
    - b.  $V/m = D$
    - c.  $DV = m$
    - d.  $D + V = m$
  22. The density of aluminum is  $2.70 \text{ g/cm}^3$ . The volume of a solid piece of Aluminum is  $1.50 \text{ cm}^3$ . Find the mass
  23. The density of pure diamond is  $3.5 \text{ g/cm}^3$ . The mass of a diamond is  $0.25 \text{ g}$ . Find the volume
  24. What statement about density is true?
    - a. two samples of a pure substance may have different densities if they are different sizes
    - b. the density of a sample depends on its location on Earth
    - c. a cylinder is always used to measure the volume
    - d. density is a physical property that remains constant for a pure substance
  25. The number of grams equal to  $0.5 \text{ kg}$  is \_\_\_\_\_
  26. In a graph the \_\_\_\_\_ variable is placed on the x axis and the \_\_\_\_\_ variable is placed on the y axis.
  27. The variable that you change in the lab is the \_\_\_\_\_ variable. The variable that responds to the change is called the \_\_\_\_\_ variable. All other things do not change and are called \_\_\_\_\_
  28. At first Dalton thought the atom was \_\_\_\_\_, then Rutherford discovered the \_\_\_\_\_ of the atom, then the subatomic particles \_\_\_\_\_, \_\_\_\_\_, and \_\_\_\_\_ were discovered
  29. The positively charged particle is the \_\_\_\_\_ found in the \_\_\_\_\_ of the atom
  30. The subatomic particle that has about the same mass as the proton, but with no electrical charge, is called a \_\_\_\_\_ and is found in the \_\_\_\_\_ of the atom
  31. The \_\_\_\_\_ is negatively charged and is found in the \_\_\_\_\_ of the atom
  32. \_\_\_\_\_ is the person credited with placing electrons on levels and \_\_\_\_\_ established the electron cloud theory.
  33. The forces that hold the particles together in the nucleus are \_\_\_\_\_ forces
  34. Isotopes are atoms of the same element that \_\_\_\_\_
  35. All isotopes of hydrogen contain \_\_\_\_\_
  36. What is the atomic number for Nitrogen? To what subatomic particle does this correspond?
  37. Zinc -66 has \_\_\_\_\_ protons and \_\_\_\_\_ neutrons
  38. What is the electron configuration for sodium? Chlorine? Neon?
  39. What is the dot notation for magnesium? Bromine? Argon?
  40. How many electrons is each element trying to achieve in the valence level?
  41. Which family is entirely radioactive?
  42. Which family will gain two electrons to complete their octet?
  43. Where are the transition metals located? What is special about their electron configuration that gives them special properties?
  44. As you move left to right across a period on the table the size (radius) \_\_\_\_\_

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**Siddappa N.Byrareddy**



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