

Chemistry Formulas



Dilution

Energy of a photon

Pressure

Average Kinetic Energy

Ideal Gas Law

Combined Gas Law

Dalton's Law of Partial

Pressures

Dalton's Law of Partial

Pressures.

Graham's Law of Effusion

Henry's Law

Freezing Point Depression

Boiling Point Elevation

Vapor Pressure Depression (Raoult's Law)

Osmotic Pressure (m)

Arrhenius Equation

Equilibrium Constant

Expressions

Equilibrium Constant

Expressions

Equilibrium Constant

Expressions

Reaction Quotient (Q)

Solubility Product Constant

 (K_{ep})

HOq & Hq

pH & pOH

pH & pOH

 $M_*V_* = M_*V_*$

Entertain = hf = hc/y

P = F/A

 $KE_{ave} = \frac{3}{2}RT$

PV = nRT

 $\frac{P_1V_1}{N_1T_1} = \frac{P_2V_2}{N_2T_2}$

P total = PA + PB+

 $P_A = \chi A P_{\text{total}}$

 $\frac{r_1}{r_2} = \sqrt{\frac{M_{min}}{M_{min}}}$

 $P_A = k_B [A]$

 $\Delta T_F = -iK_F M$

 $\Delta T_B = i K_B m$

 $P_{solu} = \chi_{solv} P^{o}_{solv}$

 $m = \ell MRT$

 $k = Ae^{-\epsilon \omega/\pi \tau}$

 $K_c = \frac{[products]}{[reactants]}$

 $K_{eq} = \frac{k_{forward}}{k_{forward}}$

 $K_P = \frac{P}{P}$

 $Q = \frac{[produstcts]}{[reactants]}$

 $K_{sp} = \frac{[products]}{[reactants]}$

 $pH = -\log[H^{-\epsilon}]$

 $pOH = -\log[OH]$

pH + pOH = 1+

pK, & pK,

pK, & pK, pK, & pK,

Neutralization Reaction

Buffers

Enthalpy of Formation

First Law of Thermodynamics

Pressure Volume Work

Calorimetry Thermal Energy

(q)

Heat Curves & Thermal Energy (g)

Heat Curves & Thermal Energy (q)

Heat Curves & Thermal Energy (q)

Entropy (S)

Bond Dissociation Energy

Gibb's Free Energy (AG)

Gibb's Free Energy (4G)

Gibb's Free Energy (4G)

Standard Cell Potential

Standard Cell Potential

Nernst Equation

Kinetics (always 1st order)

Kinetics (always 1st order)

Kinetics (always 1st order)

Nuclear Binding Energy

 $pK_a = -\log[K_a]$

 $pK_b = -\log[K_b]$

 $pK_a + pK_b = 14$

 $n_A M_A V_A = n_B M_B V_B$

 $pH = pK\alpha + \log [A -] [HA]$

 $\Delta H_f = En\Delta H^*_f(product) - En\Delta H^*_f(reactants)$

 $\Delta E = q + w$

 $w = -P\Delta V$

 $q = -C_{colorimeter}\Delta T$

 $q = mc\Delta T$

 $q = m\Delta H_{funion}$

 $q = m\Delta H_{paportration}$

 $\Delta S = \Sigma n S_{produces} - \Sigma n S_{reactions}$

 $\Delta H = \Sigma \Delta H_{restrants} - \Sigma \Delta H_{products}$

 $\Delta G^* = \Delta H^* - T \Delta S^*$

 $\Delta G = \Delta G^* + RT lnQ$

 $\Delta G^* = -RT^* ln(K_{--})$

E* = E*reduction + E*oxidation

 $E^* = E^* cathode + E^* anode$

 $E_{cell} = E^* - \frac{0.0592}{n} \log \Omega$

 $N = N_{\phi} e^{-kz}$

 $\ln N = \ln N_o - kt$

 $t_{1/2} = \frac{0.693}{k}$

 $E = \Delta mc^2$

Modern Chemistry Ch 12 Tests A And

Albert A Gayle

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