



7. What is the maximum mass of aluminum hydroxide that can be produced when 25.0 mL of 0.248 M sodium hydroxide solution are mixed with 25.0 mL of 0.189 M aluminum nitrate solution?

Find mol of each reactant

$$0.0250 \text{ L} \times 0.248 \frac{\text{mol}}{\text{L}} = 6.2 \times 10^{-3} \text{ mol NaOH}$$

$$0.0250 \text{ L} \times 0.189 \frac{\text{mol}}{\text{L}} = 4.73 \times 10^{-3} \text{ mol Al}(\text{NO}_3)_3$$

$$6.20 \times 10^{-3} \text{ mol NaOH} \times \frac{1 \text{ mol Al}(\text{OH})_3}{3 \text{ mol NaOH}} \times \frac{78.0 \text{ g Al}(\text{OH})_3}{1 \text{ mol}} =$$

$$\boxed{0.161 \text{ g Al}(\text{OH})_3}$$

need 3 times as much NaOH as $\text{Al}(\text{NO}_3)_3$ but we only have 1.3 X as much NaOH is limiting

8. Indicate which of the following are redox reactions. For those that are redox reactions, label the reactant that is the oxidizing agent and the reactant that is the reducing agent. You do not have to balance these reactions!



Redox agent agent

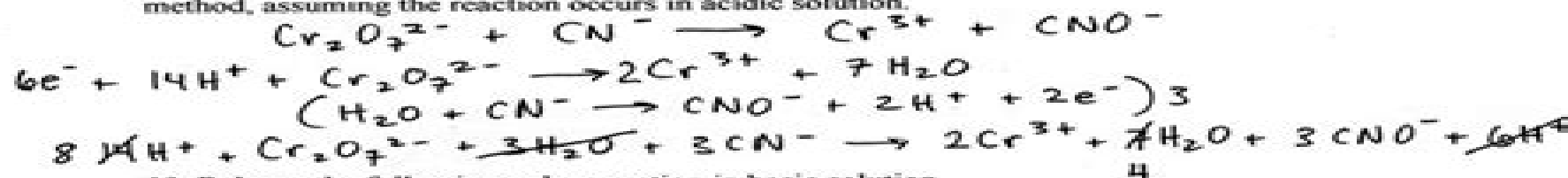


Not redox

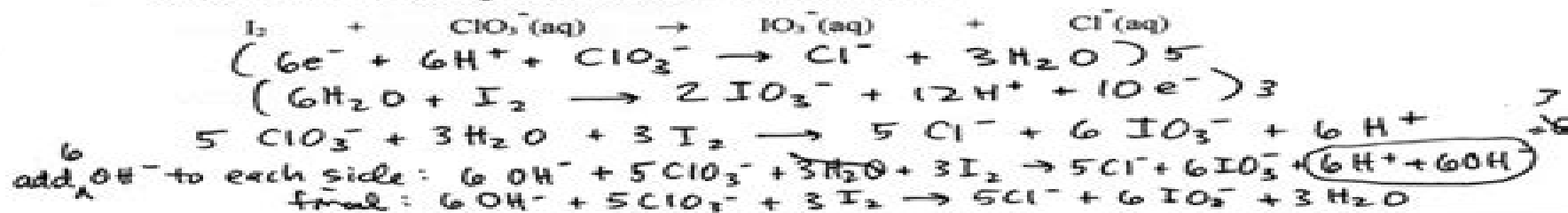


red agent ox agent

9. Aqueous dichromate ion reacts with aqueous cyanide ion (CN^-) to produce chromium (III) ion and aqueous cyanate (CNO^-) ion. Balance the equation using the half-reaction method, assuming the reaction occurs in acidic solution.



10. Balance the following redox reaction in basic solution.



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Rachel Sandford



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Table of Contents Modern Chemistry Homework 15 1 Answers

1. Understanding the eBook Modern Chemistry Homework 15 1 Answers
 - The Rise of Digital Reading Modern Chemistry Homework 15 1 Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Modern Chemistry Homework 15 1 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Modern Chemistry Homework 15 1 Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Modern Chemistry Homework 15 1 Answers
 - Personalized Recommendations
 - Modern Chemistry Homework 15 1 Answers User Reviews and Ratings
 - Modern Chemistry Homework 15 1 Answers and Bestseller Lists
5. Accessing Modern Chemistry Homework 15 1 Answers Free and Paid eBooks
 - Modern Chemistry Homework 15 1 Answers Public Domain eBooks
 - Modern Chemistry Homework 15 1 Answers eBook Subscription Services
 - Modern Chemistry Homework 15 1 Answers Budget-Friendly Options
6. Navigating Modern Chemistry Homework 15 1 Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Modern Chemistry Homework 15 1 Answers Compatibility with Devices
 - Modern Chemistry Homework 15 1 Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Modern Chemistry Homework 15 1 Answers
 - Highlighting and Note-Taking Modern Chemistry Homework 15 1 Answers
 - Interactive Elements Modern Chemistry Homework 15 1 Answers
8. Staying Engaged with Modern Chemistry Homework 15 1 Answers

- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Modern Chemistry Homework 15 1 Answers
9. Balancing eBooks and Physical Books Modern Chemistry Homework 15 1 Answers
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Modern Chemistry Homework 15 1 Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Modern Chemistry Homework 15 1 Answers
- Setting Reading Goals Modern Chemistry Homework 15 1 Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Modern Chemistry Homework 15 1 Answers
- Fact-Checking eBook Content of Modern Chemistry Homework 15 1 Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
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- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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