Dilution Calculations

What is the molarity (M) of a solution prepared by diluting 0.180L of 0.600 M HNO₃ to 0.540 L?

Prepare a table:

$$M_{1} = 0.600 M$$
 $V_{1} = 0.180 L$

$$M_2 = ?$$
 $V_2 = 0.540 L$

Solve dilution equation for unknown and enter values:

$$M_1V_1 = M_2V_2$$

$$M_2 = M_1V_1 = (0.600 \text{ M})(0.180 \text{ L}) = 0.200 \text{ M}$$
 $V_2 = 0.540 \text{ L}$

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Mocktime Publication

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