

Stoichiometry Worksheet #2 (mole-mass, mass-mole problems)



- a. If 15.0g of N_2O_4 was produced, how many moles of O_2 were required?

$$\frac{15.0\text{g N}_2\text{O}_4}{92.0\text{g N}_2\text{O}_4} \times \frac{1 \text{ mol N}_2\text{O}_4}{1 \text{ mol N}_2\text{O}_4} \times \frac{2 \text{ mol O}_2}{1 \text{ mol N}_2\text{O}_4} = 0.326 \text{ mol O}_2$$

- b. If 4.0×10^{-3} moles of oxygen reacted, how many grams of N_2 were needed?

$$\frac{4.0 \times 10^{-3} \text{ mol O}_2}{2 \text{ mol O}_2} \times \frac{1 \text{ mol N}_2}{1 \text{ mol O}_2} \times \frac{28.0\text{g N}_2}{1 \text{ mol N}_2} = 5.6 \times 10^{-2} \text{ g N}_2$$



- a. What is the mass of potassium nitrate that is produced when 2.04 moles of potassium phosphate react?

$$\frac{2.04 \text{ mol K}_3\text{PO}_4}{1 \text{ mol K}_3\text{PO}_4} \times \frac{3 \text{ mol KNO}_3}{1 \text{ mol K}_3\text{PO}_4} \times \frac{101.1 \text{ g KNO}_3}{1 \text{ mol KNO}_3} = 619\text{g KNO}_3$$

- b. If 5.80g of aluminum phosphate are formed, how many moles of aluminum nitrate reacted?

$$\frac{5.80\text{g AlPO}_4}{122.0\text{g AlPO}_4} \times \frac{1 \text{ mol AlPO}_4}{1 \text{ mol AlPO}_4} \times \frac{1 \text{ mol Al}(\text{NO}_3)_3}{1 \text{ mol AlPO}_4} = 4.75 \times 10^{-2} \text{ mol Al}(\text{NO}_3)_3$$



- a. If you have 5.50 mol of CaC_2 , how much C_2H_2 do you get?

$$\frac{5.50 \text{ mol CaC}_2}{1 \text{ mol CaC}_2} \times \frac{1 \text{ mol C}_2\text{H}_2}{1 \text{ mol CaC}_2} \times \frac{26.0\text{g C}_2\text{H}_2}{1 \text{ mol C}_2\text{H}_2} = 143\text{g C}_2\text{H}_2$$

- b. How many moles of water are needed when 65.0g of CaC_2 have reacted?

$$\frac{65.0\text{g CaC}_2}{64.1\text{g CaC}_2} \times \frac{1 \text{ mol CaC}_2}{1 \text{ mol CaC}_2} \times \frac{2 \text{ mol H}_2\text{O}}{1 \text{ mol CaC}_2} = 2.03 \text{ mol H}_2\text{O}$$

4. In photosynthesis, water reacts with carbon dioxide to give oxygen and glucose ($\text{C}_6\text{H}_{12}\text{O}_6$). Write and balance the chemical equation. How many moles of CO_2 are required to make 120.0g of glucose?



$$\frac{120.0\text{g C}_6\text{H}_{12}\text{O}_6}{180.0 \text{ g C}_6\text{H}_{12}\text{O}_6} \times \frac{1 \text{ mol C}_6\text{H}_{12}\text{O}_6}{1 \text{ mol C}_6\text{H}_{12}\text{O}_6} \times \frac{6 \text{ mol CO}_2}{1 \text{ mol C}_6\text{H}_{12}\text{O}_6} = 4.000 \text{ mol CO}_2$$

Mole Lab Answer Key

G Thomas

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the 'G Thomas' text.

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Workflow Students will read the pages in their book and then complete each section of the teacher guide. They should be encouraged to complete as many of the activities and projects as possible as well. Tests are given at regular intervals with space to record each grade.

About the Author DR DENNIS ENGLIN earned his bachelor's from Westmont College, his master of science from California State University, and his EdD from the University of Southern California. He enjoys teaching animal biology, vertebrate biology, wildlife biology, organismic biology, and astronomy at The Master's University. His professional memberships include the Creation Research Society, the American Fisheries Association, Southern California Academy of Sciences, Yellowstone Association, and Au Sable Institute of Environmental Studies.

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