

	Species	Lewis Structure	Molecular Geometry		Polar (answer yes or no and indicate polarity on Lewis drawing)
			Model drawing	VSEPR (electron geometry name and molecular structure name from Table 1)	
c.	$\text{BH}_3$	$\text{H} \cdots \text{B} \cdots \text{H}$ $\text{H}$	$\text{H}$   $\text{H}-\text{B}-\text{H}$	Trigonal Planar	no
f.	$\text{NO}_3^-$	$\text{O} \cdots \text{N} \cdots \text{O}$ $\text{O}$	$\text{O}=\text{N}^--\text{O}$	Trigonal planar Trigonal planar	no
g.	$\text{HCO}_2^-$	$\text{H} \cdots \text{C} \cdots \text{O}$ $\text{O}$	$\text{H}-\text{C}=\text{O}$ $\text{O}^-$	Trigonal planar	yes
h.	$\text{NF}_3$	$\text{F} \cdots \text{N} \cdots \text{F}$ $\text{F}$	$\text{F}-\text{N}-\text{F}$ $\text{F}$	Tetrahedral Pyramidal	yes
i.	$\text{CCl}_4$	$\text{Cl} \cdots \text{C} \cdots \text{Cl}$ $\text{Cl}$	$\text{Cl}-\text{C}-\text{Cl}$ $\text{Cl}$	Tetrahedral	yes
j.	$\text{H}_3\text{O}^+$	$\text{H} \cdots \text{O} \cdots \text{H}$ $\text{H}$	$\text{H}-\text{O}^+-\text{H}$ $\text{H}$	Tetrahedral Pyramidal	yes
k.	$\text{H}_2\text{S}$	$\text{H} \cdots \text{S} \cdots \text{H}$ $\text{S}$	$\text{H}-\text{S}-\text{H}$	Tetrahedral Bent	yes
l.	$\text{OBr}_2$	$\text{Br} \cdots \text{O} \cdots \text{Br}$	$\text{Br}-\text{O}-\text{Br}$	Tetrahedral Bent	yes

# Lewis Structure And Molecular Geometry Flinn Chemtopic Worksheet

**Arne Haaland**



## **Lewis Structure And Molecular Geometry Flinn Chemtopic Worksheet:**

**Lewis Structures and Molecular Geometry (computer File).** Jeffrey R. Buell, 1968      **Lewis Structures and Molecular Geometry** Jeffrey R. Buell, 1992      *Molecules and Models* Arne Haaland, 2008-03-06 This book provides a systematic description of the molecular structures and bonding in simple compounds of the main group elements with particular emphasis on bond distances bond energies and coordination geometries The description includes the structures of hydrogen halogen and methyl derivatives of the elements in each group some of these molecules are ionic some polar covalent The survey of molecules whose structures conform to well established trends is followed by representative examples of molecules that do not conform We also describe electron donor acceptor and hydrogen bonded complexes Chemists use models to systematize our knowledge to memorize information and to predict the structures of compounds that have not yet been studied The book provides a lucid discussion of a number of models such as the Lewis electron pair bond and the VSEPR models the spherical and polarizable ion models and molecular orbital calculations and it outlines the successes and failures of each      **Molecular Geometry** Alison Rodger, Mark Rodger, 1995 Molecular Geometry discusses topics relevant to the arrangement of atoms The book is comprised of seven chapters that tackle several areas of molecular geometry

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