

## Holt Modern Chemistry Review

### CHAPTER 6: CHEMICAL BONDING

*The following pages contain the bulk (but not all) of the information for the chapter 6 test. Focus on this content, but make sure to review class notes, activities, handouts, questions, etc. If you study this document and NOTHING else, you should at least be able to PASS the test.*  
\*\*\*\*\* Test items will be recall, examples, and/or application of this content. \*\*\*\*\*

#### OUTCOMES

- Collaborate with peer(s) to understand chemistry content (C C)
- Communicate chemistry content to teacher and peer(s) (E C)
- 6.2: Use the Lewis dot notation, illustrate the covalent bond as the sharing of electrons (F & PK)
- 6.3: List important physical properties of ionic compounds; sum of oxidation numbers = zero (F & PK)

#### 6.1: INTRODUCTION TO CHEMICAL BONDING

- **Vocabulary**
  - **chemical bond** – the attractive force that holds atoms or ions together
  - **ionic bonding** – a force that attracts electrons from one atom to another, which transforms a neutral atom into an ion
  - **covalent bonding** – a bond formed when atoms share one or more pairs of electrons
  - **nonpolar-covalent bond** – a covalent bond in which the bonding electrons are equally attracted to both bonded atoms
  - **polar** – describes a molecule in which the positive and negative charges are separated
  - **polar-covalent bond** – a covalent bond in which a pair of electrons shared by two atoms is held more closely by one atom
- **Chapter Highlights**
  - Most atoms are chemically bonded to other atoms. The three major types of chemical bonding are ionic, covalent, and metallic.
  - In general, atoms of metals bond ionically with atoms of nonmetals, atoms of metals bond metallically with each other, and atoms of nonmetals bond covalently with each other.

#### 6.2: COVALENT BONDING AND MOLECULAR COMPOUNDS

- **Vocabulary**
  - **molecule** – a group of atoms that are held together by chemical forces; a molecule is the smallest unit of matter that can exist by itself and retain all of a substance's chemical properties
  - **molecular compound** – a chemical compound whose simplest units are molecules
  - **chemical formula** – a combination of chemical symbols and numbers to represent a substance
  - **molecular formula** – a chemical formula that shows the number and kinds of atoms in a molecule, but not the arrangement of the atoms
  - **electron-dot notation** – an electron configuration notation in which only the valence electrons of an atom of a particular element are shown, indicated by dots placed around the element's symbol
  - **Lewis structure** – a structural formula in which electrons are represented by dots; dot pairs or dashes between two atomic symbols represent pairs in covalent bonds.  
See examples at right for LEWIS STRUCTURES of NF<sub>3</sub> and OCl<sub>2</sub>
  - **structural formula** – a formula that indicates the location of the atoms, groups, or ions relative to one another in a molecule and that indicates the number and location of chemical bonds
  - **single bond** – a covalent bond in which two atoms share one pair of electrons
  - **multiple bond** – a bond in which the atoms share more than one pair of electrons, such as a double bond or a triple bond
  - **resonance** – the bonding in molecules or ions that cannot be correctly represented by a single Lewis structure
- **Chapter Highlights**
  - Atoms in molecules are joined by covalent bonds. In a covalent bond, two atoms share one or more pairs of electrons.



# Modern Chemistry Chapter Review Chemical Bonding

**Clemens Wendtner**



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## **Table of Contents Modern Chemistry Chapter Review Chemical Bonding**

1. Understanding the eBook Modern Chemistry Chapter Review Chemical Bonding
  - The Rise of Digital Reading Modern Chemistry Chapter Review Chemical Bonding
  - Advantages of eBooks Over Traditional Books
2. Identifying Modern Chemistry Chapter Review Chemical Bonding
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Modern Chemistry Chapter Review Chemical Bonding
  - User-Friendly Interface
4. Exploring eBook Recommendations from Modern Chemistry Chapter Review Chemical Bonding
  - Personalized Recommendations
  - Modern Chemistry Chapter Review Chemical Bonding User Reviews and Ratings
  - Modern Chemistry Chapter Review Chemical Bonding and Bestseller Lists
5. Accessing Modern Chemistry Chapter Review Chemical Bonding Free and Paid eBooks
  - Modern Chemistry Chapter Review Chemical Bonding Public Domain eBooks
  - Modern Chemistry Chapter Review Chemical Bonding eBook Subscription Services
  - Modern Chemistry Chapter Review Chemical Bonding Budget-Friendly Options

6. Navigating Modern Chemistry Chapter Review Chemical Bonding eBook Formats
  - ePub, PDF, MOBI, and More
  - Modern Chemistry Chapter Review Chemical Bonding Compatibility with Devices
  - Modern Chemistry Chapter Review Chemical Bonding Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Modern Chemistry Chapter Review Chemical Bonding
  - Highlighting and Note-Taking Modern Chemistry Chapter Review Chemical Bonding
  - Interactive Elements Modern Chemistry Chapter Review Chemical Bonding
8. Staying Engaged with Modern Chemistry Chapter Review Chemical Bonding
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Modern Chemistry Chapter Review Chemical Bonding
9. Balancing eBooks and Physical Books Modern Chemistry Chapter Review Chemical Bonding
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Modern Chemistry Chapter Review Chemical Bonding
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Modern Chemistry Chapter Review Chemical Bonding
  - Setting Reading Goals Modern Chemistry Chapter Review Chemical Bonding
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Modern Chemistry Chapter Review Chemical Bonding
  - Fact-Checking eBook Content of Modern Chemistry Chapter Review Chemical Bonding
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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