
CHEMISTRY- 6.8 Kinetic Molecular Theory: A Model for Gases- STUDY GUIDE WITH 100% CORRECT SOLUTIONS

Kinetic Molecular Theory - ✓ ✓ Model explaining gas behavior through particle motion.

Gas Particles - ✓ ✓ Collection of molecules or atoms in constant motion.

Elastic Collision - ✓ ✓ Collision where total kinetic energy is conserved.

Negligible Volume - ✓ ✓ Particles occupy insignificant space compared to distances.

Average Kinetic Energy - ✓ ✓ Proportional to temperature in kelvins for gas particles.

Thermal Energy - ✓ ✓ Energy due to temperature affecting particle motion.

Velocity Distribution - ✓ ✓ Variation in particle speeds at a given temperature.

Argon Gas Example - ✓ ✓ 0.01% of volume occupied by argon atoms at STP.

Atomic Radius of Argon - ✓ ✓ 97 pm; average distance between atoms is 3.3 nm.

Collision Mechanics - ✓ ✓ Particles do not exert forces between collisions.

Billiard Ball Analogy - ✓ ✓ Elastic collisions compared to gas particle interactions.

Modern Chemistry Gases Study Guide

RM Cervero



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