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CHAPTER 6 REVIEW

Chemical Bonding

MIXED REVIEW

SHORT ANSWER Answer the following questions in the space provided.

1. Name the type of energy that is a measure of strength for each of the following types of bonds:

a. ionic bond

b. covalent bond

e. metallic bond

Use the electronegativity values shown in Figure 5-20, on page 151 of the text, to determine whether each of the following bonds is nonpolar covalent, polar covalent, or ionic.

a. H-F

______ b. Na—Cl _____ e. H—C

_____ c. H—O _____ f. H—N

3. How is a hydrogen bond different from an ionic or covalent bond?

 H₂S and H₂O have similar structures and their central atoms belong to the same group. Yet H₂S is a gas at room temperature and H-O is a liquid. Use bonding principles to explain why this is true.

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Modern Chemistry Chapter 6 Review

J Dewey

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