

# AP Chemistry Pacing Guide

<b>Unit 4: Chemical Reactions</b> <ul style="list-style-type: none"> <li>Introduction for Reactions</li> <li>Net Ionic Equations</li> <li>Representations of Reactions</li> <li>Physical and Chemical Changes</li> <li>Stoichiometry</li> <li>Introduction to Titration</li> <li>Types of Chemical Reactions</li> <li>Introduction to Acid-Base Reactions</li> <li>Oxidation-Reduction (Redox) Reactions</li> </ul>	<b>BIG IDEA 1:</b> Scale, Proportion, and Quantity <b>BIG IDEA 3:</b> Transformations	<ol style="list-style-type: none"> <li>I can identify evidence of chemical and physical changes in matter.</li> <li>I can represent changes in matter with a balanced chemical or net ionic equation: a. For physical changes; b. For given information about the identity of the reactants and/or product.</li> <li>For ions in a given chemical reaction I can represent a given chemical reaction or physical process with a consistent particulate model.</li> <li>I can explain the relationship between macroscopic characteristics and bond interactions for: a. Chemical processes; b. Physical processes.</li> <li>I can explain changes in the amounts of reactants and products based on the balanced reaction equation for a chemical process.</li> <li>I can identify the equivalence point in a titration based on the amounts of the titrant and analyte, assuming the titration reaction goes to completion.</li> <li>I can identify a reaction as acid-base, oxidation-reduction, or precipitation.</li> <li>I can identify species as Brønsted-Lowry acids, bases, and/or conjugate acid-base pairs, based on proton-transfer involving those species.</li> <li>I can represent a balanced redox reaction equation using half-reactions.</li> </ol>	10 days
<b>Unit 5: Kinetics</b> <ul style="list-style-type: none"> <li>Reaction Rates</li> <li>Introduction to Rate Law</li> <li>Concentration Changes Over Time</li> <li>Elementary Reactions</li> <li>Collision Model</li> <li>Reaction Energy Profile</li> <li>Introduction to Reaction Mechanism</li> <li>Mechanism Reaction Mechanism and Rate Law</li> <li>Steady-State</li> </ul>	<b>BIG IDEA 3:</b> Transformations <b>BIG IDEA 6:</b> Energy	<ol style="list-style-type: none"> <li>I can explain the relationship between the rate of a chemical reaction and experimental parameters.</li> <li>I can represent experimental data with a consistent rate law expression.</li> <li>I can identify the rate law expression of a chemical reaction using data that show how the concentrations of reaction species change over time.</li> <li>I can represent an elementary reaction as a rate law expression using stoichiometry.</li> <li>I can explain the relationship between the rate of an elementary reaction and the frequency, energy, and orientation of molecular collisions.</li> <li>I can represent the activation energy and overall energy change in an elementary reaction using a reaction energy profile.</li> <li>I can identify the components of a reaction mechanism.</li> <li>I can identify the rate law for a reaction from a mechanism in which the first step is rate limiting.</li> <li>I can identify the rate law for a reaction from a mechanism in which the</li> </ol>	10 days

# Modern Chemistry Pacing Guide

**KJ Lindholm-Leary**



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## **Table of Contents Modern Chemistry Pacing Guide**

1. Understanding the eBook Modern Chemistry Pacing Guide
  - The Rise of Digital Reading Modern Chemistry Pacing Guide
  - Advantages of eBooks Over Traditional Books
2. Identifying Modern Chemistry Pacing Guide
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Modern Chemistry Pacing Guide
  - User-Friendly Interface
4. Exploring eBook Recommendations from Modern Chemistry Pacing Guide
  - Personalized Recommendations
  - Modern Chemistry Pacing Guide User Reviews and Ratings
  - Modern Chemistry Pacing Guide and Bestseller Lists

5. Accessing Modern Chemistry Pacing Guide Free and Paid eBooks
  - Modern Chemistry Pacing Guide Public Domain eBooks
  - Modern Chemistry Pacing Guide eBook Subscription Services
  - Modern Chemistry Pacing Guide Budget-Friendly Options
6. Navigating Modern Chemistry Pacing Guide eBook Formats
  - ePub, PDF, MOBI, and More
  - Modern Chemistry Pacing Guide Compatibility with Devices
  - Modern Chemistry Pacing Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Modern Chemistry Pacing Guide
  - Highlighting and Note-Taking Modern Chemistry Pacing Guide
  - Interactive Elements Modern Chemistry Pacing Guide
8. Staying Engaged with Modern Chemistry Pacing Guide
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Modern Chemistry Pacing Guide
9. Balancing eBooks and Physical Books Modern Chemistry Pacing Guide
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Modern Chemistry Pacing Guide
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Modern Chemistry Pacing Guide
  - Setting Reading Goals Modern Chemistry Pacing Guide
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Modern Chemistry Pacing Guide
  - Fact-Checking eBook Content of Modern Chemistry Pacing Guide
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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