

The Mole

Section 10.1 Measuring Matter

pages 320–324

Practice Problems

pages 323–324

1. Zinc (Zn) is used to form a corrosion-inhibiting surface on galvanized steel. Determine the number of Zn atoms in 2.50 mol of Zn.

$$2.50 \text{ mol-Zn} \times \frac{6.02 \times 10^{23} \text{ atoms Zn}}{1 \text{ mol-Zn}} \\ = 1.51 \times 10^{24} \text{ atoms of Zn}$$

2. Calculate the number of molecules in 11.5 mol of water (H₂O).

$$11.5 \text{ mol-H}_2\text{O} \times \frac{6.02 \times 10^{23} \text{ molecules H}_2\text{O}}{1 \text{ mol-H}_2\text{O}} \\ = 6.92 \times 10^{24} \text{ molecules of H}_2\text{O}$$

3. Silver nitrate (AgNO₃) is used to make several different silver halides used in photographic films. How many formula units of AgNO₃ are there in 3.25 mol AgNO₃?

$$3.25 \text{ mol-AgNO}_3 \times \frac{6.02 \times 10^{23} \text{ formula units AgNO}_3}{1 \text{ mol-AgNO}_3} \\ = 1.96 \times 10^{24} \text{ formula units of AgNO}_3$$

4. **Challenge** Calculate the number of oxygen atoms in 5.0 mol of oxygen molecules. Oxygen is a diatomic molecule, O₂.

$$5.00 \text{ mol-O}_2 \times \frac{6.02 \times 10^{23} \text{ molecules-O}_2}{1 \text{ mol-O}_2} \\ \times \frac{2 \text{ O atoms}}{1 \text{ molecule-O}_2} = 6.02 \times 10^{24} \text{ atoms O}$$

5. How many moles contain each of the following?

a. 5.75×10^{24} atoms Al

$$5.75 \times 10^{24} \text{ atoms-Al} \times \frac{1 \text{ mol Al}}{6.02 \times 10^{23} \text{ atoms-Al}} \\ = 9.55 \text{ mol Al}$$

b. 2.50×10^{20} atoms Fe

$$2.50 \times 10^{20} \text{ atoms-Fe} \times \frac{1 \text{ mol Fe}}{6.02 \times 10^{23} \text{ atoms-Fe}} \\ = 4.15 \times 10^{-4} \text{ mol Fe}$$

6. **Challenge** Identify the representative particle for each formula and convert the given number of representative particles to moles.

a. 3.75×10^{24} CO₂

The representative particle is a molecule.

$$3.75 \times 10^{24} \text{ molecules-CO}_2 \\ \times \frac{1 \text{ mol CO}_2}{6.02 \times 10^{23} \text{ molecules-CO}_2} = 6.23 \text{ mol CO}_2$$

b. 3.58×10^{23} ZnCl₂

The representative particle is a formula unit.

$$3.58 \times 10^{23} \text{ formula-units-ZnCl}_2 \\ \times \frac{1 \text{ mol ZnCl}_2}{6.02 \times 10^{23} \text{ formula-units-ZnCl}_2} \\ = 0.595 \text{ mol ZnCl}_2$$

Section 10.1 Assessment

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7. **Explain** why chemists use the mole.

Chemists use the mole because it is a convenient way of knowing how many representative particles are in a sample.

8. **State** the mathematical relationship between Avogadro's number and 1 mol.

One mole contains 6.02×10^{23} representative particles.

9. **List** the conversion factors used to convert between particles and moles.

$$\frac{6.02 \times 10^{23} \text{ representative particles}}{1 \text{ mol}} \\ \frac{1 \text{ mol}}{6.02 \times 10^{23} \text{ representative particles}}$$

Mole Solution Manual

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Papadantonakis**

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Physical Chemistry, Student Solutions Manual Robert J. Silbey, Robert A. Alberty, Mounji G. Bawendi, George A. Papadantonakis, 2021-12-06 The Fifth Edition of the Student Solutions Manual Physical Chemistry delivers the answers to all four types of problems offered in Physical Chemistry as well as the computer problems The Solutions Manual provides full worked out solutions for the exercises that can be solved with a hand held calculator and Mathematica solutions for all 170 problems that require a personal computer This book also facilitates digital access to all Mathematica answers at www.wiley.com/go/silbey/physicalchemistry5e

Chemistry, Student Solutions Manual James N. Spencer, George M. Bodner, Lyman H. Rickard, 2002-07-25 The second edition of Spencer's Chemistry Structure and Dynamics has been the most successful reform project published for the General Chemistry course The authors have revised the text by building on the recommendations of the ACS's Task Force on the General Chemistry Curriculum and suggestions from the adopters of the first edition This innovative text provides a fifteen chapter introduction to the fundamental concepts of Chemistry A collection of additional topics at the end of each chapter allow instructors to supplement and tailor their courses according to individual need Three major themes link the content of the book the process of science the relationship between molecular structure and physical chemical properties and the relationship between the microscopic and macroscopic levels

Study Guide/Selected Solutions Manual Julie R. Frentrop, 2002-06 Contains a brief overview of every chapter review of skills self tests and the answers and detailed solutions to all odd numbered end of chapter problems in the text book

Chemistry, Student Solutions Manual James E. Brady, Joel W. Russell, John R. Holum, 2000-01-26 The images on the cover call attention to the relationship between macro observations and the intimate structure of chemical substances and the changes both chemical and physical that they undergo

Fireworks One of the ingredients is phosphorus a molecular form of which is believed to consist of linked tetrahedra of phosphorus atoms The chemical reaction of phosphorus with oxygen is partly responsible for the spectacular show of light

Carbon The element is found in several forms including the familiar diamond and another recently discovered sooty substance that consists of soccer ball shaped molecules often referred to as buckeyballs

Diamond is not the most stable form of carbon and is created from other forms of carbon at high temperatures and pressures deep within the earth

Acetylene torch Cutting steel is possible because of the intense heat generated by the chemical reaction of acetylene with oxygen a reaction between molecules of C_2H_2 and O_2 to give CO_2 and H_2O

Hot air balloon The air that helps it rise is heated by the combustion of molecules of propane each composed of three carbon and eight hydrogen atoms

Stormy weather The evaporation of water serves to store energy provided by the sun Subsequent condensation of the water vapor releases this energy and is the basis of all the weather systems on our planet

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Study Guide and Solutions Manual for Seager/Slabaugh's Chemistry for Today Seager/Slabaugh, 2004 The fifth edition of the Study Guide and Student Solutions Manual has been updated to reflect all of the changes to the text This ancillary tests the student on the learning objectives in each chapter and provides answers to all of the even numbered end of chapter exercises New additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application

Student Solutions Manual for Chang's Chemistry Philip C. Keller, Raymond Chang, Jill L. Keller, 1988

Student Solutions Manual Deborah Wiegand, 1999-07

Student Study Guide and Solutions Manual to Accompany General, Organic, and Biochemistry Katherine J. Denniston, Denniston Katherine, 2006

Physical Chemistry, Solutions Manual Robert A. Alberty, Robert J. Silbey, 1996-08-20 This book provides thorough coverage of physical chemistry It demonstrates the power and limits of thermodynamics with a more systematic treatment of the second law and more focus on entropy It also covers current topics in physical chemistry and shows how physical chemistry relates to daily life Includes many current applications such as lasers

Instructor's Solutions Manual to Accompany Introductory Chemistry Daniele Peters, 1997

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Solutions Manual and Study Guide to Accompany Introduction to Organic Chemistry, 4th Ed Paul A. Bartlett, Judith G. Koch, 1992

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