

ChemThink: Molecular Shapes

Objective:




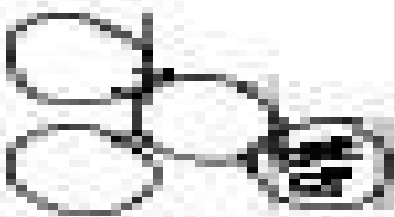

Lesson 8

Electrons

Page

Electron pairs

Understanding that as the number of molecules that have lone pairs on an atom, the electron pairs will repel each other and the shape repulsion will affect the shape of the molecule.

Number of e- pairs on atom	Draw an example	Bond Angles	Shape	E.g. Lewis dot representation
2		180°	Linear	$H-H$
3		120°	Trigonal Planar	
4		109°	Tetrahedral	

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C Cleary



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