#### 14: Chemical Bonding Theories

#### Key Chemistry Terms

- I omic Bend: Bond formed from electrostatic attraction between ions (charged atoms)—formed from metals with nonmetals.
- Covalent Bond: A bond formed between two nonmetals that involves shared electrons.
- Potar covalent bond: Bond formed between two nonmetals—uneven sharing of electrons.
- Metallic bond: Formed between metal atoms—electrons are pooled between the network of atoms.
- Potar bond: Bond between two atoms with a great difference in electronegativities—uneven sharing of electrons resulting in a slightly positive region and a slightly regative region of the bond.
- Electronegativity: An atom's "pull" on electrons shared with another atom.
- Potar molecule: Molecule where the polar bonds do not cancel each other out in 3D orientation.
- I somers: Molecules with the same chemical formula, but different bonding structure.
- Resonance: Two compounds with multiple bonds in different locations.
- Valence Bond Theory: Overlap of atomic orbitals form bond.
- Sigma (e) bonds first bond between two atoms formed from head on overlap of orbitals.
- Pf (g) bond: 2<sup>nd</sup> or 3<sup>nd</sup> bond between two atoms formed from overlap of parallel p orbitals.
- Hybridization: Atomic orbitals hybridize into orbitals of the same energy for all sigma bonds of the atom.

#### Characteristics of bond types

Bond type	Happens. between	Electrons are
Formic:	Metal & non-metal	Transferred
Cowalent	Non-metals	Scharced
Polar Covalent	Non-metals	Shared unevenly
Meta Sic	Metals	p-pooled

Common characteristics:

- Lonic: High melting points, most dissolve in water, conduct electricity when dissolved in water, brittle.
- Covalent: Low melting points, most do not dissolve in water, do not conduct electricity when dissolved in water.
- Polar covalent: Medium melting points, some dissolve in water, do not conduct electricity when dissolved in water.
- Metallic: Soft, conduct heat and electricity, do not dissolve in water.

#### Bond Polarity

When nonmetals bond covalently with a large difference in electronegativity.

- Absolute value of differences:
  - o 0 0.4 m covalent
  - a 0.5 1.4 = polar covalent
  - a 1.5 4 = lonic

Symbolized with an arrow pointing towards the more electronegative element and a crossed tail by the less electronegative element.

#### Cocamode:

C - N Electronegativity difference of 0.5 = polar bond



#### I somers and Resonance

Exiomiteir ascarmiphess:

Riesionanice examples:

#### Sigma and Pi bonds

How many sigma and pi bonds are in the following?

6 sigma bonds & 3 pi bonds

Hybridiza	rtion
-----------	-------

# of sigma bonds	Hybridization	Formed from
2	Sp	1 s and 1 p
3	50	1 s and 2 p
4	500	I s and 3 p

How to Use This Cheat Sheet. These are the keys related this topic. Try to read through it carefully twice then rewrite it out on a tilank sheet of paper. Review it again before the exams.

# **Molecular Structure And Bonding Cheat Sheet**

**R Barnett** 

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