

# CHEMICAL KINETICS

Pg-1

## Rate of Reaction :

$$\text{Rate} = \frac{\text{Change in Conc or Pressure}}{\text{Time taken}}$$

$$\text{Unit} \rightarrow \text{mol L}^{-1} \text{T}^{-1} \text{ or atm T}^{-1}$$

\* ROR can never be negative. (-ve)

## Rate of Reaction. (A $\rightarrow$ B)

### Average Rate

$$\text{Rate} = -\frac{\Delta(A)}{\Delta t} = \frac{\Delta(B)}{\Delta t}$$



$$R_{\text{avg}} = \left( \frac{A_i - A_f}{t_2 - t_1} \right)$$

### Instantaneous Rate

$$\text{Rate} = -\frac{d(A)}{dt} = \frac{d(B)}{dt}$$



$$\text{Rate} = \text{slope} = \left| \frac{dy}{dx} \right|$$



$$\text{Rate (ROR)} = -\frac{d(\text{N}_2)}{dt} = -\frac{d(\text{H}_2)}{3dt} = \frac{1}{2} \frac{d(\text{NH}_3)}{dt}$$

$$\text{Rate of Disappearance} = -\frac{d(\text{N}_2)}{dt} = -\frac{d(\text{H}_2)}{dt}$$

$$\text{Rate of Appearance} = \frac{d(\text{NH}_3)}{dt} \text{ (Product)}$$

ROR & RODA does not include the St. Coeff.



$$\text{ROR} = -\frac{\text{RODA}}{a} = \frac{\text{RODB}}{b} = \frac{\text{RODC}}{c} = \frac{\text{RODD}}{d}$$

$$\text{ROR} = -\frac{1}{a} \frac{d(A)}{dt} = -\frac{1}{b} \frac{d(B)}{dt} = \frac{1}{c} \frac{d(C)}{dt} = \frac{1}{d} \frac{d(D)}{dt}$$

~ Differential Rate Law

Q. The rate of the reaction:  $2\text{N}_2\text{O}_5 \rightarrow 4\text{NO}_2 + \text{O}_2$   
Can be written in three ways  $\rightarrow$

$$-\frac{d(\text{N}_2\text{O}_5)}{dt} = K [\text{N}_2\text{O}_5] ; \frac{d(\text{NO}_2)}{dt} = k' [\text{N}_2\text{O}_5]$$

$$\frac{d(\text{O}_2)}{dt} = K'' [\text{N}_2\text{O}_5]$$

$$\text{Soln. ROR} = -\frac{1}{2} \frac{d(\text{N}_2\text{O}_5)}{dt} = \frac{1}{4} \frac{d(\text{NO}_2)}{dt} = \frac{d(\text{O}_2)}{dt}$$

$$= -\frac{K [\text{N}_2\text{O}_5]}{2} = \frac{K' [\text{N}_2\text{O}_5]}{4} = \frac{K'' [\text{N}_2\text{O}_5]}{1}$$

$$\text{ROR} = K' = 2K \quad K'' = \frac{K}{2}$$

$$Q = 0.3 + \log \left( -\frac{d(A)}{dt} \right) = 0.18 + \log \left( \frac{d(B)}{dt} \right)$$

write the reaction.

$$0.3 + \log \left( -\frac{d(A)}{dt} \right) = 0.18 + \log \left( \frac{d(B)}{dt} \right)$$

$$0.3 + \log \left( -\frac{d(A)}{dt} \right) = 0.18 + \log \left( \frac{d(B)}{dt} \right)$$



## Molecularity :

Δ Can be integer (1, 2 or 3) but it cannot be zero, negative or fractional.

Δ For elementary reaction: Molecularity is equal to its order.

Δ Maximum value of molecularity is 3. (or order)

Δ For complex rxn molecularity of each step of mechanism is defined separately.  $\therefore$  molecularity cannot be defined for whole complex reaction at once.

Δ Molecularity of a reaction may or may not be equal to order of reaction.

Δ Molecularity can not be determined by balanced chemical rxn and can only be determined from simple reaction.

# Manual Solution Of Chemical Kinetics

**Charles G. Hill**



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Solutions Manual: Sm Chemical Kinetics and React Dyn Paul L. Houston, 2001      *Chemical Engineering Kinetics* Joseph Mauck Smith, C. Y. Cha, 1971      **Chemical Engineering Kinetics** Joseph Mauk Smith, 1983      Chemical Kinetics, **An Introduction to Chemical Engineering Kinetics and Reactor Desing** Charles G. Hill, 1977      *Catalog of Copyright Entries. Third Series* Library of Congress. Copyright Office, 1976      *Invariant Manifolds for Physical and Chemical Kinetics* Alexander N. Gorban, Iliya V. Karlin, 2005-02-01 By bringing together various ideas and methods for extracting the slow manifolds the authors show that it is possible to establish a more macroscopic description in nonequilibrium systems The book treats slowness as stability A unifying geometrical viewpoint of the thermodynamics of slow and fast motion enables the development of reduction techniques both analytical and numerical Examples considered in the book range from the Boltzmann kinetic equation and hydrodynamics to the Fokker Planck equations of polymer dynamics and models of chemical kinetics describing oxidation reactions Special chapters are devoted to model reduction in classical statistical dynamics natural selection and exact solutions for slow hydrodynamic manifolds The book will be a major reference source for both theoretical and applied model reduction Intended primarily as a postgraduate level text in nonequilibrium kinetics and model reduction it will also be valuable to PhD students and researchers in applied mathematics physics and various fields of engineering      Physical Chemistry Peter Atkins, Julio de Paula, 2002 New edition of the overwhelmingly favorite text for the physical chemistry course      *Chemical Kinetics and Dynamics* Jeffrey I. Steinfeld, Joseph Salvadore Francisco, William L. Hase, 1999 This text presents a balanced presentation of the macroscopic view of empirical kinetics and the microscopic molecular viewpoint of chemical dynamics This second edition includes the latest information as well as new topics such as heterogeneous reactions in atmospheric chemistry reactant product imaging and molecular dynamics of H<sub>2</sub>      **Monthly Catalog of United States Government Publications, Cumulative Index** United States. Superintendent of Documents, 1970      Chemical Kinetics and Reaction Mechanisms James H. Espenson, 1981      *Enzyme Kinetics: Catalysis and Control* Daniel L. Purich, 2010-06-16 Far more than a comprehensive treatise on initial rate and fast reaction kinetics this one of a kind desk reference places enzyme science in the fuller context of the organic inorganic and physical chemical processes occurring within enzyme active sites Drawing on 2600 references Enzyme Kinetics Catalysis Control develops all the kinetic tools needed to define enzyme catalysis spanning the entire spectrum from the basics of chemical kinetics and practical advice on rate measurement to the very latest work on single molecule kinetics and mechanoenzyme force generation while also focusing on the persuasive power of kinetic isotope effects the design of high potency drugs and the behavior of regulatory enzymes Historical analysis of kinetic principles including advanced enzyme science Provides both theoretical and practical measurements tools Coverage of single molecular kinetics Examination of force generation mechanisms Discussion of organic and inorganic enzyme reactions      *Kinetics of Catalytic Reactions--Solutions Manual* M.

Albert Vannice, 2008-03-27 This manual of solutions to the problems in Kinetics of Catalytic Reactions has been prepared to assist those who use this book in a teaching function. However, these solutions should also benefit those outside the classroom who want to apply the principles and concepts that are discussed in the book. By studying and observing the approaches used in solving these problems, it is very likely that similar applications can be envisioned in different kinetic problems that the investigator might face. Thus, the availability of these solutions is a good learning tool for everyone. Additional details and insight about the solutions provided can be obtained by reading the cited references. I have tried to eliminate all errors, both conceptual and typographical, in these solutions; however, the probability is high that I have not succeeded completely. Should any errors of commission or omission be found, I would greatly appreciate being informed. I can be reached at this email address: mavche@engr.psu.edu or mail can be sent to me at 107 Fenske Laboratory, Department of Chemical Engineering, The Pennsylvania State University, University Park, PA 16802.

Albert Vannice v Contents Preface v Solutions to Problems Chapter 3 Catalyst Characterization      **Monthly Catalog of United States Government Publications** United States.

Superintendent of Documents, 1966      **U.S. Environmental Protection Agency Library System Book Catalog Holdings as of July 1973** United States. Environmental Protection Agency. Library Systems Branch, 1974      Journal of the American Chemical Society American Chemical Society, 1928 Proceedings of the Society are included in v 1 59 1879 1937

**American Druggist**, 1923      **Scientific and Technical Books in Print**, 1972      Scientific and Technical Books and Serials in Print, 1984      *Reference Catalogue of Current Literature*, 1920

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