

Limiting Reactant Worksheet #2

1. Given the following reaction: (Balance the equation first!)



- If you start with 14.8 g of C_3H_8 and 3.44 g of O_2 , determine the limiting reactant.
- Determine the number of moles of carbon dioxide produced.
- Determine the number of grams of H_2O produced.
- Determine the number of grams of excess reactant left.

2. Given the following equation:



- If 10.0 g of $\text{Al}_2(\text{SO}_4)_3$ is reacted with 10.0 g of NaOH , determine the limiting reactant.
- Determine the number of moles of $\text{Al}(\text{OH})_3$ produced.
- Determine the number of grams of Na_2SO_4 produced.
- Determine the number of grams of excess reactant left over in the reaction.

3. Given the following equation:



- If 25.4 g of Al_2O_3 is reacted with 10.2 g of Fe , determine the limiting reactant.
- Determine the number of moles of Al produced.
- Determine the number of grams of Fe_3O_4 produced.
- Determine the number of grams of excess reactant left over in the reaction.

4. Consider the reaction



- 80.0 grams of I_2O_5 , reacts with 28.0 grams of carbon monoxide, CO . Determine the mass of iodine I_2 , which could be produced?
- If, in the above situation, only 0.160 moles, of iodine, I_2 was produced.
 - What mass of iodine was produced?
 - What percentage yield of iodine was produced?

5. Zinc and sulfur react to form zinc sulfide according to the equation



If 25.0 g of zinc and 30.0 g of sulfur are mixed,

- Which chemical is the limiting reactant?
- How many grams of ZnS will be formed?
- How many grams of the excess reactant will remain after the reaction is over?

6. Which element is in excess when 3.00 grams of Mg is ignited in 2.20 grams of pure oxygen? What mass is in excess? What mass of MgO is formed?

7. How many grams of Al_2S_3 are formed when 5.00 grams of Al is heated with 10.0 grams S ?

Limiting Reagent Practice 2 Answers

Guillaume Favre



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McGraw-Hill's PCAT George J. Hademenos, Shaun Murphree, Kathy A. Zahler, Mark Whitener, Jennifer M. Warner, 2008-10-31 Complete preparation for the more than 30 000 students taking the PCAT each year The number of Pharmacy College Admission Test PCAT test takers has doubled since 2001 and that number continues to grow McGraw Hill's PCAT provides comprehensive review of every test topic while preparing you for the new format of the test that was instituted in June 2007 With sample tests test taking strategies and intricate scientific illustrations you ll have all the preparation you need to perform at your very best on test day **General Chemistry** Raymond Chang, 2003 Publisher

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