

## 2. Limiting and Excess Reactants

The **limiting reactant** is defined as the species in a chemical reaction that **theoretically** would be the first to be completely consumed if the reaction were to proceed to completion according to the chemical equation.

*- even if the reaction does not proceed to completion!*

All of the other reactant/s are called **excess reactants**.

For Example:



If 1 g mol of  $\text{C}_7\text{H}_{16}$  and 12 g mol of  $\text{O}_2$  are mixed so as to react,

# Limiting And Excess Reactants Chembridge Pogil

**Lifelique**



### **Limiting And Excess Reactants Chembridge Pogil:**

**Limiting Reactant and Percent** Lifeliqe,2019 This lesson plan covers analyzing chemical reactions in order to determine limiting reactants and excess reactants including calculating the amount of excess reactant calculate the theoretical yield of a reaction and calculating the percent yield of a reaction

Delve into the emotional tapestry woven by Crafted by in **Limiting And Excess Reactants Chembridge Pogil** . This ebook, available for download in a PDF format ( \*), is more than just words on a page; it's a journey of connection and profound emotion. Immerse yourself in narratives that tug at your heartstrings. Download now to experience the pulse of each page and let your emotions run wild.

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