

Name _____

Date _____

Limiting Reactant Lab

Pre-lab Questions: Use the background information in the lesson and the video to help you answer these questions before the lab.

1. What is a limiting reactant? (10pts)

A limiting reactant in a chemical reaction determines how much product will be produced.

2. What is an excess reactant? (10pts)

An excess reactant in a chemical reaction is present in an amount greater than what will be used to complete the reaction

3. What is the purpose of this experiment? (10pts)

to see the reaction of aluminum foil when placed in an aqueous solution of copper (II) chloride to produce copper and aluminum chloride. The purpose is to determine which of the reactants will limit the production of copper

Procedures: Use the video in the lesson to help you complete the data table and post-lab questions below.

Data & Observation Table (12 points)	
Mass of Al (g)	0.12 g
Volume of CuCl ₂ (mL)	50.9 mL
Color of CuCl ₂ solution before reaction	green
Color of Al before reaction	Silver
Color of CuCl ₂ solution after reaction	dark cloudy green
Color of Al after the reaction	silver and brown

4. Balance the equation below: (8 pts)



Limiting Reactant Quick Lab Answers

Kenneth W. Raymond



Limiting Reactant Quick Lab Answers:

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With this modular laboratory program students build skills using important chemical concepts and techniques to the point where they are able to design a solution to a scenario drawn from a professional environment The scenarios are drawn from the lives of people who work with chemistry every day ranging from field ecologists to chemical engineers and include many health professionals as well *Laboratory Safety for Chemistry Students* Robert H. Hill, Jr., David C. Finster, 2011-09-21 this substantial and engaging text offers a wealth of practical in every sense of the word advice Every undergraduate laboratory and ideally every undergraduate chemist should have a copy of what is by some distance the best book I have seen on safety in the undergraduate laboratory Chemistry World March 2011 Laboratory Safety for Chemistry Students is uniquely designed to accompany students throughout their four year undergraduate education and beyond progressively teaching them the skills and knowledge they need to learn their science and stay safe while working in any lab This new principles based approach treats lab safety as a distinct essential discipline of chemistry enabling you to instill and sustain a culture of safety among students As students progress through the text they ll learn about laboratory and chemical hazards about routes of exposure about ways to manage these hazards and about handling common laboratory emergencies Most importantly they ll learn that it is very possible to safely use hazardous chemicals in the laboratory by applying safety principles that prevent and minimize exposures Continuously Reinforces and Builds Safety Knowledge and Safety Culture Each of the book s eight chapters is organized into three tiers of sections with a variety of topics suited to beginning intermediate and advanced course levels This enables your students to gather relevant safety information as they advance in their lab work In some cases individual topics are presented more than once progressively building knowledge with new information that s appropriate at different levels A Better Easier Way to Teach and Learn Lab Safety We all know that safety is of the utmost importance however instructors continue to struggle with finding ways to incorporate safety into their curricula Laboratory Safety for Chemistry Students is the ideal solution Each section can be treated as a pre lab assignment

enabling you to easily incorporate lab safety into all your lab courses without building in additional teaching time Sections begin with a preview a quote and a brief description of a laboratory incident that illustrates the importance of the topic References at the end of each section guide your students to the latest print and web resources Students will also find Chemical Connections that illustrate how chemical principles apply to laboratory safety and Special Topics that amplify certain sections by exploring additional relevant safety issues Visit the companion site at <http://userpages.wittenberg.edu/dfinster/LSCS>

Improving Student Comprehension in Chemistry Laboratories Tracy Lynn Haroff, 2006 *Student Study Guide and Solutions Manual to accompany General Organic and Biological Chemistry, 1e* Kenneth W. Raymond, 2005-10-07 Finally readers have a shorter less intimidating introduction to general organic and biological chemistry Not only is Raymond's text concise it also takes an integrated approach to presenting important topics in a way that makes the material easier to understand In this approach similarities can be exploited and concepts reinforced The result is that readers see the strong connections that exist between these three branches of chemistry

Laboratory Experiments for General Chemistry Harold R. Hunt, Toby F. Block, George M. McKelvy, 1998 *SourceBook Version 2.1*, 1998 *Chemistry* John S. Phillips, Cheryl Wistrom, 2000

Teaching Science for Understanding James Joseph Gallagher, 2007 Offers middle and high school science teachers practical advice on how they can teach their students key concepts while building their understanding of the subject through various levels of learning activities

Instructors Manual to Lab Manual Ralph Petrucci, William Harwood, Geoffrey Herring, 2001 *Scientific and Technical Aerospace Reports*, 1985 Lists citations with abstracts for aerospace related reports obtained from world wide sources and announces documents that have recently been entered into the NASA Scientific and Technical Information Database

Laboratory Manual to Accompany Introductory Chemistry Uno Kask, Otto W. Nitz, Martin G. Ondrus, Tammy J. Melton, 1993-07 This fifth edition of this laboratory manual emphasizes safety in the lab and discusses equipment requirements in the apparatus section at the beginning of each experiment It also features a revised art programme and explains the rationale for each experiment

Improving Student Comprehension of Stoichiometric Concepts Connie Lynn Bannick Kemner, 2007

Introduction to Chemical Principles Howard Stephen Stoker, 1986

Semimicro Experiments in Organic Chemistry Bruce Nelson Campbell, Ernest Ray Kline, 1964

Chemistry Neil D. Jespersen, Alison Hyslop, 2021-12-03 *Chemistry The Molecular Nature of Matter* 8th Edition continues to focus on the intimate relationship that exists between structure at the atomic molecular level and the observable macroscopic properties of matter Key revisions in this edition focus on three areas The deliberate inclusion of more updated real world examples that relate common real world student experiences to the science of chemistry Simultaneously examples and questions have been updated to align them with career concepts relevant to the environmental engineering biological pharmaceutical and medical sciences Providing students with transferable skills with a focus on integrating metacognition and three dimensional learning into the text When students know what they know they are better able to learn and

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