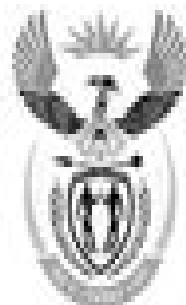


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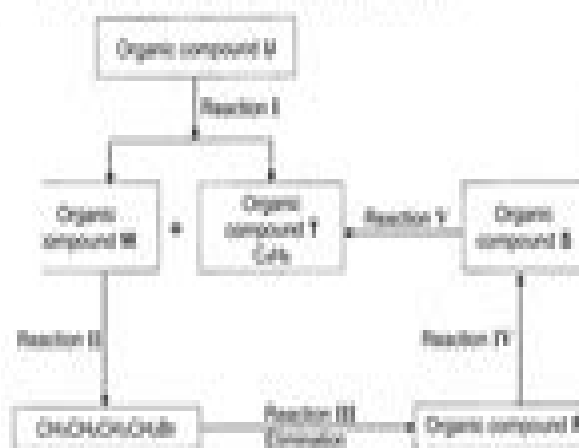
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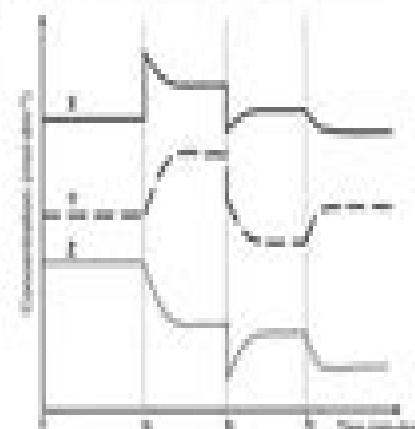
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Diagram on page 9 is redrawn below for easy reference.



Changes to the conditions of equilibrium are made at different times. The graph shows the results obtained. X, Y and Z represent the gases in the above reaction.



2.1.1 Define for both chemical equilibrium

Use the graph to answer the questions below.

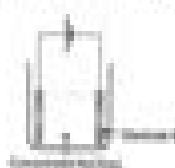
The inorganic reagents shown below are available for reactions IV and Y.

Bu	$\text{H}_2\text{SO}_4(\text{conc})$	$\text{NaOH}(\text{conc})$	HBr	H_2
----	--------------------------------------	----------------------------	-----	--------------

4.5 Write down the TYPE of elimination in reaction III.

4.6 Compounds R and T are positional isomers.

4.7 Identify the reagent used in reaction IV. (1 mark)



4.8 Write the half-cell reaction for the anode. (1 mark)

4.9 Write the overall cell reaction. (1 mark)

4.10 Write down the name of the reaction shown in the diagram below. (1 mark)



4.11 Write down the name of the reaction shown in the diagram below. (1 mark)

Reaction I: $\text{Pb} + \text{PbO}_2 + \text{H}_2\text{SO}_4 \rightarrow \text{PbSO}_4 + \text{H}_2\text{O}$

4.12 Write down the name of the reaction shown in the diagram below. (1 mark)

4.13 Write down the name of the reaction shown in the diagram below. (1 mark)

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