

## Modern Analytical Chemistry Worksheet

1. What is the relationship between absorbance and transmittance? (3 points)

Transmittance is the fraction of radiation (from the source) that passes through the sample. It is given by the formula  $T = P/P_0$ . Meanwhile, absorbance is the intensity of radiation absorbed by the sample and it is the negative logarithm of transmittance ( $A = -\log T$ ).

2. A solution containing the complex formed between Bi (III) and thiourea has a molar absorptivity of  $9.32 \times 10^3 \text{ L/cm}\cdot\text{mol}$  at 470nm.

(a) What is the absorbance of a  $5.67 \times 10^{-5} \text{ M}$  solution of the complex at 470 nm in a 1.00 cm-cell?

$$A = \epsilon bc$$

Given:

$$\epsilon = 9.32 \times 10^3 \text{ L/cm}\cdot\text{mol}$$

$$b = 1.00 \text{ cm}$$

$$c = 5.67 \times 10^{-5} \text{ M}$$

$$A = ?$$

$$\text{Solution: } A = (9.32 \times 10^3) \times (1.00 \text{ cm}) \times (5.67 \times 10^{-5}) = 0.528$$

(b) What is the percent transmittance of the solution described in (a)

$$A = -\log T$$

Given:

$$A = 0.528$$

$$\%T = ?$$

Solution:

$$-\log T = 0.528$$

$$T = (10^{0.528})^{-1} = 0.296483139$$

$$\%T = 0.296483139 \times 100\% = 29.6\%$$

(c) What is the molar concentration of the complex in a solution that has the absorbance described in (a) when measured at 470 nm in a 2.50 cm cell?

$$A = \epsilon bc$$

Given:

$$A = 0.528$$

$$\epsilon = 9.32 \times 10^3 \text{ L/cm}\cdot\text{mol}$$

$$b = 2.50 \text{ cm}$$

$$c = ?$$

Solution:

$$c = \frac{0.528}{(9.32 \times 10^3 \text{ L/cm}\cdot\text{mol})(2.50 \text{ cm})} = 2.27 \times 10^{-5} \text{ M}$$

# Modern Chemistry Review Mixed Review

**William Thomas Stead**



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