

Key

## Molecular Geometry

How can molecular shapes be predicted using the VSEPR theory?

### Why?

When you draw a Lewis structure for a molecule on paper, you are making a two-dimensional representation of the atoms. In reality however, molecules are not flat—they are *three-dimensional*. The true shape of a molecule is important because it determines many physical and chemical properties for the substance. In this activity you will learn how to predict molecular shapes.

### Model 1 – Lewis Structures

Lewis Structures		3-D Molecular Shape
1. $\text{H}_2\text{CO}$	3 electron domains (3 bonding, 0 nonbonding)	trigonal planar
2. $\text{BeF}_2$	2 electron domains (2 bonding, 0 nonbonding)	Linear
3. $\text{CH}_4$	4 electron domains (4 bonding, 0 nonbonding)	Tetrahedral
4. $\text{NH}_3$	4 electron domains (3 bonding, 1 nonbonding)	pyramidal
5. $\text{H}_2\text{O}$	4 electron domains (2 bonding, 2 nonbonding)	Bent
6. $\text{CO}_2$	2 electron domains (2 bonding, 0 nonbonding)	linear

Lone pair = ..

# Molecular Geometry Pogil Answer Key

**Brendan G. Carr**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

## **Molecular Geometry Pogil Answer Key:**

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