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(a) The symbol equation for the reaction can be represented as shown in Fig. 3.1.

Fig. 3.1

Table 3.1 shows some bond energies.

Bond	H-H	N-N
Bond energy in kJ/mol	435	193

Use the bond energies in Table 3.1 and  $\Delta H = -91 \text{ kJ/mol}$  to calculate the bond energy of an N-H bond in kJ/mol.

Use the following steps.

- Calculate the energy needed to break bonds in the reactants.
 
$$\begin{aligned}
 & (4 \times \text{H-H}) + 1(\text{N-N}) \\
 & 4 \times 435 + 1(193)
 \end{aligned}$$

3150 kJ
- Calculate the energy released when bonds form in the products.
 
$$\begin{aligned}
 & \text{Bond broken} - \text{Bond formed} = -91 \\
 & 3150 - x = -91 \\
 & x = 3150 + 91
 \end{aligned}$$

3240 kJ
- Calculate the energy of per mole bond.
 
$$\begin{aligned}
 & 6 \times (\text{N-H}) = 3240 \\
 & \text{N-H} = \frac{3240}{6}
 \end{aligned}$$

540 kJ/mol

# Mark Scheme 0620 63 M J 13

**ML Yell**



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**The Christian Advocate** ,1901

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