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(a) The symbol equation for the reaction can be represented as shown in Fig. 3.1.

Fig. 3.1

Table 3.1 shows some bond energies.

bond	kJ/mol	kJ/mol
bond energy in kJ/mol	440	435

Use the bond energies in Table 3.1 and $\Delta H = -2140 \text{ kJ/mol}$ to calculate the bond energy of one C-H bond in kJ/mol .

Use the following steps.

- Calculate the energy needed to break bonds in the reactants.

$$\begin{aligned}
 & (500) + 3(440) \\
 & 440 + 3(440) \quad \quad \quad 2160 \text{ kJ}
 \end{aligned}$$
- Calculate the energy released when bonds form in the products.

$$\begin{aligned}
 & \text{bond broken} - \text{bond formed} = -90 \\
 & 2160 - x = -90 \\
 & x = 2250 \text{ kJ}
 \end{aligned}$$
- Calculate the energy of C-H bond.

$$\begin{aligned}
 & 6 \times \text{C-H} = 2140 \\
 & \text{C-H} = \frac{2140}{6} \quad \quad \quad 357 \text{ kJ/mol}
 \end{aligned}$$

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