Annotated Solution 2018 USNCO Local Exam

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1 Sobutions

1. We can use Assignfus's constant, and then do must dissensional analysis.

$$225 \pm O_2 = \frac{1 \text{ and } O_2}{30.0 \pm O_2} = \frac{2 \text{ and } O}{1 \text{ and } O_2} = \frac{6.02 \pm 20^{12} \text{ atoms}}{\text{ned } O} = \frac{9.42 \pm 20^{21} \text{ atoms}}{10.02 \pm 20^{21} \text{ atoms}}$$

Thus, the surrect in C1.

2. Decause the same number of moles must remain before and after dilution, we can set up the following equation:

$$M_1N_1 - M_2N_2$$

(1.50 M)($N_1 - 0.300 M$)(0.500 L)
(0.50 M)($N_1 - 0.100 M - L$
 $N_1 - 0.100 L$

This is 1000 set. or 4

3. We first need to write the reaction in order to perform the stoichiometry. The reaction is no follows:

Now we can do the stoichingstry.

$$10.00 \pm \mathrm{Cu} \times \frac{1 \bmod \mathrm{Cu}}{60.05 \pm \mathrm{Cu}} \times \frac{1 \bmod \mathrm{H_2O}}{2 \bmod \mathrm{Cu}} \times \frac{19.00 \pm \mathrm{H_2O}}{1 \bmod \mathrm{H_2O}} = \begin{bmatrix} 1.417 \pm \mathrm{H_2O} \end{bmatrix}$$

These, the success in [41].

 Set a standard assessed of energies at an arbitrary 100.0 g. Thus, there are 48.61 g Cu. 19.02 g Au, and 32.52 g S. We can divide by the moles masses of each of the compounds.

$$48.41 \text{ g Cu} \times \frac{1 \text{ mod Cu}}{628.55 \text{ g Cu}} = 0.7618 \text{ mod}$$

 $19.00 \text{ g As} \times \frac{1 \text{ mod As}}{74.722 \text{ g As}} = 0.2039 \text{ mod}$
 $32.62 \text{ g S} \times \frac{1 \text{ mod S}}{20.06 \text{ g S}} = 1.046 \text{ mod}$

Dividing by the lowest number, then there will be the ratio of $Cu : As : S \hookrightarrow B : 1 : 4$. Thus, the formula is $Cu_1AsS_1[]$ or [C].

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