13.2

THE NATURE OF LIQUIDS

Section Review

Objectives

- Identify factors that determine physical properties of a liquid.
- Define avaporation in terms of kinetic energy
- Describe the equilibrium between a liquid and its vapor
- · Identify the conditions under which boiling occurs

Vocabulary

- vaporization
- evaporation.
- vapor pressure

- boiling point
- normal boiling point

Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

Liquids are much ______ than gases. Liquids and solids

are known as _____ states of matter. The conversion of a liquid

to a gas or vapor is called ______ When a liquid that is not
 ______ thanges to a gas, the process is called evaporation. A

Liquid evaporates faster when heated; however, evaporation itself

is a ______ process. When a partially filled container of liquid is

sealed, some of the particles at the ______ of the liquid vaporize.

These particles collide with the walls of the container, producing

a force called _____ The vapor pressure of a liquid can be

determined by a device called a _____ A liquid boils when its

______ equals the external pressure. The normal boiling point of a

Part B True-False

Classify each of these statements as always true, AT; sometimes true, ST; or never true, NT.

 Particles in a liquid don't have enough kinetic energy to overcome the attractive forces between them and vaporize.

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Lei Huang

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