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> Coursebook answers

Chapter 16



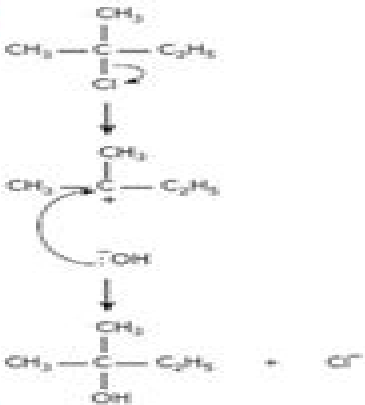
Science in context

Discuss as a group how the information about the use of halogenoalkanes demonstrates how some scientific developments can have both good and bad effects on society and the environment.

- Having read the Science in Context passage, ask learners to shut their Coursebook. Then conduct a quick activity asking individual learners to give a specific example of a beneficial use of halogenoalkanes or a problematic use, until the class runs out of suggestions.
- You might then ask for a class vote, show of hands, on whether the halogenoalkanes have benefitted or hindered human progress in the last century.

Self-assessment questions

- propene and chlorine
 - OH^- ions are negatively charged so are attracted more strongly than neutral water molecules to the partially positively charged carbon atoms in halogenoalkanes.
 - The hydrolysis of halogenoalkanes produces halide ions; the rate of their formation can be monitored by using silver nitrate solution. The silver halide precipitates make the reaction mixture cloudy. The ionic equations for the formation of the precipitates are:
 $\text{Ag}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{AgCl}(\text{s})$
 $\text{Ag}^+(\text{aq}) + \text{Br}^-(\text{aq}) \rightarrow \text{AgBr}(\text{s})$
 $\text{Ag}^+(\text{aq}) + \text{I}^-(\text{aq}) \rightarrow \text{AgI}(\text{s})$
 - Both ammonia and amines contain a nitrogen atom with a lone pair of electrons that is available to donate.
 - tripropylamine $(\text{CH}_3\text{CH}_2\text{CH}_2)_3\text{N}$

- 
 - 
 - 
- $\text{CH}_3\text{CHBrCH}_3 + \text{NaOH}(\text{ethanol}) \rightarrow \text{CH}_2=\text{CHCH}_3 + \text{H}_2\text{O} + \text{NaBr}$
 - propene
 - $\text{BrCH}_2\text{CH}_2\text{CH}_3 + \text{NaOH}(\text{ethanol}) \rightarrow \text{CH}_2=\text{CHCH}_3 + \text{H}_2\text{O} + \text{NaBr}$
 (1-bromopropane)

Mastering Chemistry Answers Chapter 16

L Towner

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