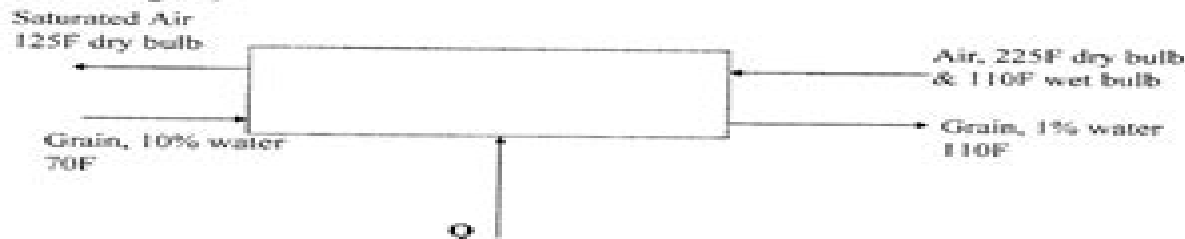


- Air at 45°C and 10% relative humidity is to be humidified adiabatically to 60% relative humidity. Use the psychrometric chart to obtain the following numbers:
  - (5 points) The dry bulb and adiabatic saturation temperature of the final air.
  - (5 points) The dew point of the final air
  - (5 points) The absolute humidity and humid volume of final air.
  - (5 points) The rate at which water must be added to humidify 15m<sup>3</sup>/min of the entering air.
- (20 points) Two aqueous sulfuric acid solutions containing 20.0 wt% H<sub>2</sub>SO<sub>4</sub> (specific gravity=1.139) and 60.0 wt % H<sub>2</sub>SO<sub>4</sub> (specific gravity=1.498) are mixed to form a 4.00 molar solution (specific gravity=1.213). Assuming 100 kg of 20% solution as the basis calculate:
  - (10 points) the mass fraction of sulfuric acid in the product solution.
  - (10 points) the feed ratio (liters of 20% solution/liter of 60% solution).
- (30 points) A rotary dryer operating at atmospheric pressure dries 10 tons/day of wet grain at 70°F, from a moisture content of 10% to 1% moisture. The air flow is countercurrent to the flow of grain, enters at 225°F dry bulb and 110°F wet-bulb temperature, and leaves at 125°F dry-bulb (see figure).



Determine: (a) the humidity of entering and leaving air if the latter is saturated, (b) the water removal in pounds per hour, (c) the daily product output in pounds per day, and (d) the heat input to the dryer. Assume that there is no heat loss from the dryer, that the grain is discharged at 110°F, and that its specific heat is 0.18.

- (30 points) A fuel gas containing 85 mole % methane and the balance ethane is burned completely with pure oxygen at 25 °C at 1 atm. Suppose the reactor is continuous. Take a basis of calculation of 1 mol/s of the fuel gas, combusted with stoichiometric amount of oxygen.
  - (10 points) Please calculate the rate at which heat must be transferred from the reactor if the products are brought back down to 25 °C.
  - (15 points) Please calculate the final temperature of the combustion gas if the process is under adiabatic condition at constant pressure 1 atm.
  - (5 points) Please calculate the final volume of the combustion gas if the process is under adiabatic condition at constant pressure 1 atm. If the volume of the reactor can only be constant and the pressure can change, what do you expect to happen to the reactor?

# Lifesciencep1 Final Exam 2014

**David Baud**



## **Lifesciencep1 Final Exam 2014:**

*Teaching Psychology* Douglas A. Bernstein, Sue Frantz, Stephen L. Chew, 2019-11-08 This thoroughly revised third edition of *Teaching Psychology* synthesizes the latest pedagogical research on effective teaching and translates it into recommendations for classroom application. It also takes into account the many changes in the teaching landscape that have taken place in recent years. Covering key topics such as planning a course, choosing teaching methods, assimilating technology, and the integration of teaching into the rest of your academic life, this book also includes an abundance of supportive supplementary content to guide and inform new teachers. This content will also benefit seasoned teachers who wish to reevaluate their current teaching practices and explore new teaching ideas and techniques. Presenting a comprehensive and cutting edge teaching guide for psychology teachers, this book is a vital resource for those who are training psychology instructors or undertaking a teaching psychology course. It is also a useful text for more experienced faculty looking to update their current teaching practices.

**Government Reports Announcements & Index**, 1979

**Harper's Weekly** John Bonner, George William Curtis, Henry Mills Alden, Samuel Stillman Conant, Montgomery Schuyler, John Foord, Richard Harding Davis, Carl Schurz, Henry Loomis Nelson, John Kendrick Bangs, George Brinton McClellan Harvey, Norman Hapgood, 1875

**SET Life Science: Solved Exam Questions** Kailash Choudhary, D. Sondge, R.P. Saran, N. Soni, 2017-12-01 The present book *SET Life Science Solved Papers* is specially developed for the aspirants of SET Life Sciences Examinations. This book includes previous solved papers SET Life Science papers of Maharashtra Andhra Pradesh Karnataka Tamil Nadu Kerala Gujarat and Rajasthan. Main objective of this book is to develop confidence among the candidates appearing for SET examination in the field of Life Sciences. Both fundamental and practical aspects of the subject have been covered by solved questions. This book meets the challenging requirements of CSIR NET GATE IARI BARC and Ph D entrance of various Indian universities.

*Prentice-Hall Life Science*, 1986

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**Workflow:**

- Step 1 Teacher leads Discussion Questions
- Step 2 Student watches DVD if applicable
- Step 3 Student reads chapter
- Step 4 Student is given and completes Chapter Objectives after reading the chapter
- Step 5 Teacher administers Chapter Test
- Step 6 Teacher administers Sectional Exams where indicated
- Step 7 Teacher administers Comprehensive Final Exams where indicated

**Prentice Hall Life Science** Elizabeth Coolidge-Stoltz, Michael J. Padilla, Ioannis Miaoulis, Martha Cyr, 2009

*Prentice-Hall Life Science* Charles R. Coble, Dale R. Rice, 1986

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**Science** , 1987

*The Life Science* Peter Brian Medawar, Jean Medawar, 1977

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