

JANUARY ... DOES IT BRING YOU JOY?

- 1 January is the time for renewal on many fronts. It's the time of the year to re-evaluate everything. If something or someone doesn't bring you joy, it's pretty pointless carrying on with it. If we want the best, we must do our best. We must keep our family space safe and secure so that we are thriving as a people and as a nation. 5
- 2 In order to live their best lives, people are spreading themselves thinner and thinner with multiple streams of income. In the process we earn more but live less, we lose our focus of what is important.
- 3 With social media, keeping up with other Instagrammers are having a devastating effect on self-perception. Everything appears to be real perfect, but the fabric of our society is frazzled. Use January to take stock and start navigating new ways of doing things. Instead of multiple streams of sole sessions, travel more, and connect with people and places that bring you joy. 10

[Adapted from *The Juice*, January 2020]

QUESTIONS: TEXT F

- 5.1 Refer to line 1.
Provide a synonymous phrase for 'on many fronts'. (1)
- 5.2 What is the function of the hyphen in 're-evaluate' (line 2)? (1)
- 5.3 Rewrite the following sentence in the passive voice.
'If something or someone doesn't bring you joy, it's pretty pointless carrying on with it' (lines 2–3). (1)
- 5.4 Remove the redundancy in paragraph 1. (1)

Memorandum Of English Paper 2 Final Exam

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monoxide co determine the mass of iodine I_2 which could be produced b if in the above situation only 0.160 moles of iodine I_2 was produced i what mass of iodine was produced

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$$Cu + 2 AgNO_3 \rightarrow Cu(NO_3)_2 + 2 Ag$$

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web start with 8.3 mole H_2S and multiply it with 2 mole H_2O over 2 mole H_2S from the equation $8.3 \text{ mole } H_2S \times \frac{2 \text{ mole } H_2O}{2 \text{ mole } H_2S} = 8.3 \text{ mole } H_2O$
 $8.3 \text{ mole } H_2O \times 18 \text{ g } H_2O / \text{mole } H_2O = 149.4 \text{ g } H_2O$ is your theoretical yield so percent yield $\frac{137.1 \text{ g } H_2O}{149.4 \text{ g } H_2O} \times 100 = 91.77\%$
 problems 6 8 are similar to problem 5

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you have 0.216232 grams mole 21.3 grams 0.66 moles 32 g m

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calculated mass of the product expected $\times 100$ in this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ 4 Al 3 O_2 2 Al_2O_3 b determine the theoretical yield of Al_2O_3 160.0

g Al $\times 1 \text{ mol Al} \times 2 \text{ mol Al}_2\text{O}_3 \times 101.964 \text{ g Al}_2\text{O}_3 / 302.3 \text{ g Al}_2\text{O}_3 = 26.982 \text{ g Al}_2\text{O}_3$ 4 mol Al 1 mol Al_2O_3 c determine the percent yield

yield actual $\times 100$ 260.0 g $\times 100$ 86.01 theoretical 302.3 g

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web 7 Na 2 C 2 O 4 is the limiting reactant percent yield 86.6 8 only four molecules can be made 9 this amount cannot be

weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12

C is the limiting reagent 4 33 g of H 2 are left over 13

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from a set of reactants calculate how much product will be produced from the limiting reagent calculate how much reactant s

remains when the reaction is complete

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web the percent yield is calculated as follows $\text{percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$ percent

yield theoretical $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100$ based on this definition we would expect a percent yield

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CH_3OH can be produced by the following reaction $2\text{H}_2 + \text{CO} \rightarrow \text{CH}_3\text{OH}$ calculate the theoretical yield of CH_3OH if 68.5 g of CO is

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