

Moorpark College
Chemistry 1A Mock Exam #1 **ANSWER KEY**
Professor Torres

1. A chemistry student has the idea of removing a decorative stone sphere with a radius of 0.289 m from the front of a campus building. If the density of the stone is 3.52 g/cm^3 , what is the mass of the sphere?

$$V_{\text{sphere}} = \frac{4}{3} \pi r^3$$

$$\text{Volume} = \frac{4}{3} \pi r^3 = \frac{4}{3} \times \pi \times (28.9 \text{ cm})^3 = 1.01 \times 10^5 \text{ cm}^3$$

$$\text{Density} = \text{mass/volume} \Rightarrow \text{Mass} = \text{Density} \times \text{Volume}$$

$$\text{Mass} = (3.52 \text{ g/cm}^3)(1.01 \times 10^5 \text{ cm}^3) = \boxed{3.56 \times 10^5 \text{ g}}$$

2. Determine the mass (in grams) of 1.00 gal of cyclohexane. The density of cyclohexane at 25°C is 0.7781 g/mL . [NOTE: $1 \text{ L} = 1.057 \text{ qt}$ and $1 \text{ gal} = 4 \text{ qt}$]

$$\begin{aligned} \text{Mass (g)} &= 1.00 \text{ gal} \times \frac{4 \text{ qt}}{1 \text{ gal}} \times \frac{1 \text{ L}}{1.057 \text{ qt}} \times \frac{1000 \text{ mL}}{1 \text{ L}} \times \frac{0.7781 \text{ g}}{\text{mL}} \\ &= \boxed{2.94 \times 10^3 \text{ g cyclohexane}} \end{aligned}$$

3. CFC's, or chlorofluorocarbons, are damaging pollutants to the earth's atmosphere that are known as greenhouse gases. Above 30,000 feet, the atmosphere has a density of $1.2 \times 10^{-6} \text{ g/cm}^3$. Determine the mass of CFCl_3 (a CFC) contained in $7.2 \times 10^6 \text{ cm}^3$ air if the CFCl_3 concentration is 50. ppm (i.e. one gram of CFCl_3 in one million grams of air).

$$\begin{aligned} \text{grams CFCl}_3 &= \frac{50 \text{ g CFCl}_3}{10^6 \text{ g air}} \times \frac{1.2 \times 10^{-6} \text{ g air}}{\text{cm}^3} \times 7.2 \times 10^6 \text{ cm}^3 \text{ air} \\ &= \boxed{4.32 \times 10^{-1} \text{ g CFCl}_3} \end{aligned}$$

4. The element potassium has two naturally occurring isotopes: ^{39}K and ^{41}K with a mass of 40.962 amu. Potassium consists of 6.88% ^{41}K and has an average atomic mass of 39.101 amu. Calculate the mass of ^{39}K .

$$\begin{aligned} x &= \text{mass of } ^{39}\text{K} \\ 39.101 &= 40.962(0.0688) + 0.9312x \\ x &= \boxed{38.964 \text{ amu}} \end{aligned}$$

5. Nitrogen has two naturally occurring isotopes, ^{14}N and ^{15}N , whose atomic masses are 14.0031 and 15.0001, respectively. The atomic mass of nitrogen is 14.0067. Calculate the percentage abundance of each isotope of nitrogen.

$$\begin{aligned} x &= \% \text{ abundance of } ^{14}\text{N}; 1 - x = \% \text{ abundance } ^{15}\text{N} \\ 14.0067 &= 14.0031x + 15.0001(1 - x) \\ x &= 0.9964 \text{ or } \boxed{99.64\% \text{ } ^{14}\text{N}; \text{ therefore, } 0.36\% \text{ } ^{15}\text{N}} \end{aligned}$$

Moorpark Chemistry Answer Sheets

Joacim Rocklöv



Moorpark Chemistry Answer Sheets:

Moorpark Chemistry Answer Sheets Book Review: Unveiling the Magic of Language

In a digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "**Moorpark Chemistry Answer Sheets**," written by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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