

### QUESTION 1: FACTORISE

- 1.1 Factorise the following expression completely:

$$x\left(\frac{2x^2}{x} - \frac{3xy}{x}\right) + y\left(\frac{2xy}{y} - \frac{3y^2}{y}\right) \quad (3)$$

- 1.2 When  $f(x) = 3ax^3 + 2x^2 - 7x + 3$  is divided by  $2x + 1$  the remainder is  $-\frac{a}{2}$ .

Determine the value or values of  $a$ . (5)

- 1.3 Simplify the following expression:

$$\frac{b}{a+b} + \frac{a}{a-b} + \frac{2ab}{b^2 - a^2} \quad (6)$$

[14]

### QUESTION 2: FACTORISE – AUG 18

- 2.1 Simplify the following algebraic fraction:

$$\frac{x^2 - 4}{3x^2 + 4x - 4} \div \frac{x^2 - 4x + 4}{2x^2 - 11x + 14} \times \frac{3x - 2}{6x - 21} \quad (5)$$

- 2.2 When  $f(x)$  is divided by  $2x - 3$  the remainder is 4.

Determine the value of  $p$  if  $f(x) = 2x^3 - px^2 + 6x - 5$ . (4)

- 2.3 Factorise the following expression completely:  
 $ax + 9by - 3ay - 3bx$

(3)

- 2.4 Simplify the following expression:

$$\frac{x + 9x^{\frac{1}{2}}}{x + 7x^{\frac{1}{2}} - 18} \div \left(1 - \frac{2}{x^{\frac{1}{2}}}\right)^{-1} \quad (5)$$

[17]

### QUESTION 3: FACTORISE

- 3.1 Factorise the following as far as possible in prime factors:

3.1.1  $5x^4 - 5y^4$

3.1.2  $15x^2y^2 - 25xy^3 - 10y^4$

(2 × 3) (6)

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**L Manion**



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## **Table of Contents Ma1210 College Mathematics I Answers**

1. Understanding the eBook Ma1210 College Mathematics I Answers
  - The Rise of Digital Reading Ma1210 College Mathematics I Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Ma1210 College Mathematics I Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Ma1210 College Mathematics I Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Ma1210 College Mathematics I Answers
  - Personalized Recommendations
  - Ma1210 College Mathematics I Answers User Reviews and Ratings
  - Ma1210 College Mathematics I Answers and Bestseller Lists

5. Accessing Ma1210 College Mathematics I Answers Free and Paid eBooks
  - Ma1210 College Mathematics I Answers Public Domain eBooks
  - Ma1210 College Mathematics I Answers eBook Subscription Services
  - Ma1210 College Mathematics I Answers Budget-Friendly Options
6. Navigating Ma1210 College Mathematics I Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Ma1210 College Mathematics I Answers Compatibility with Devices
  - Ma1210 College Mathematics I Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Ma1210 College Mathematics I Answers
  - Highlighting and Note-Taking Ma1210 College Mathematics I Answers
  - Interactive Elements Ma1210 College Mathematics I Answers
8. Staying Engaged with Ma1210 College Mathematics I Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Ma1210 College Mathematics I Answers
9. Balancing eBooks and Physical Books Ma1210 College Mathematics I Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Ma1210 College Mathematics I Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Ma1210 College Mathematics I Answers
  - Setting Reading Goals Ma1210 College Mathematics I Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ma1210 College Mathematics I Answers
  - Fact-Checking eBook Content of Ma1210 College Mathematics I Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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web elements in the second row of the periodic table place their electrons in the 2n shell as well as the 1n shell for instance lithium li has three electrons two fill the 1 s orbital and the third is placed in the 2 s orbital giving an electron configuration of 1 s 2 2 s 1

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web electron configurations and trends answer these discussion questions given the following part of an electron configuration 3s<sup>2</sup> explain which part refers to the number of electrons the energy level and the sublevel what is the maximum number of electrons that can occupy a single orbital state and explain hund s rule

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web one more point needs to be emphasized about the relationship between electron configuration and the periodic table the atoms of elements in the same vertical column of the table have similar electron configurations for example consider the alkaline earth elements group iia

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web the electron removed during the ionization of beryllium he 2s 2 is an s electron whereas the electron removed during the ionization of boron he 2s 2 2p 1 is a p electron this results in a lower first ionization energy for boron even though its nuclear charge is greater by one proton

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web predictions from reliable sources have been used for these elements grayed out electron numbers indicate subshells filled to their maximum bracketed noble gas symbols on the left represent inner configurations that are the same in each period written out these are he 2 helium 1s 2 ne 10 neon 1s 2 2s 2 2p 6

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web q1 what is the electron configuration of the element potassium answer the atomic number of potassium is 19 a neutral potassium element has 19 electrons its electron configuration is 1s 2 2s 2 2p 6 3s 2 3p 6 4s 1 or simply put ar 4s 1 q2 what is

the total number of electrons that can be filled in s p and d subshells

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