

## Chemistry Unit 7 Review

### 701 – Inquiry Lab

1. Define...
  - a. Direct relationship: Two variables increase/decrease together
  - b. Indirect relationship: One variable increases, the other decreases
2. Identify whether the listed variables have a direct or indirect relationship to each other.
  - a. Moles and volume: Direct
  - b. Temperature and pressure: Direct
  - c. Pressure and volume: Indirect
  - d. Moles and pressure: Direct
3. Using the Ideal Gas Law equation, list 3 ways you could increase the volume of a balloon.  
Increase the moles, increase the temperature, decrease the pressure (more so, the pressure drops as a result of greater volume)
4. Using the Ideal Gas Law equation, list 3 ways you could decrease the pressure within a syringe.  
Decrease the moles, decrease the temperature, increase the volume
5. If the number of moles in a container is doubled, by how much should the pressure increase (assuming everything else is held constant)?  
The pressure should also double

### 702 – Gas Laws Equation

1. What equation should you use when you are given just one set of variables?  
 $PV = nRT$
2. You have a 4.5 L canister of  $O_2$  that you store in a room that is  $20\text{ }^{\circ}\text{C}$ . The pressure of this tank is 13000 kPa. How many moles of  $O_2$  are in this canister?  
24 moles

# Modeling Chemistry U7 Test V

**Jicheng Xie**



## **Modeling Chemistry U7 Test V:**

**Proceedings** ,1989      **Energy Research Abstracts** ,1982      **Technical Abstract Bulletin** ,      **Restructuring the Geothermal Industry** Geothermal Resources Council,Geothermal Resources Council. Annual Meeting,1994      Readers' Guide to Periodical Literature Anna Lorraine Guthrie,Bertha Tannehill,Neltje Marie Tannehill Shimer,1925 An author subject index to selected general interest periodicals of reference value in libraries      **Materials Performance** ,1993      Ceramic Abstracts American Ceramic Society,1926      Corrosion ,1988-07 Issues include special section called Corrosion abstracts  
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