

## ChemThink: Ionic Formulas

Noble Gases never form \_\_\_\_\_.

Atoms react with each other to become \_\_\_\_\_.

Atoms form ions that have the same number of \_\_\_\_\_ as noble gases. This fact can be used to predict the charge of an \_\_\_\_\_.

In order to form an ionic compound we will need to have one atom that can \_\_\_\_\_ electrons (\_\_\_\_\_ \_\_\_\_\_) and one atom that can \_\_\_\_\_ electrons (\_\_\_\_\_ \_\_\_\_\_.)

When sodium loses an electron what happens to the size of the blue atom?

When chlorine gains an electron what happens to the size of the green atom?

When oxygen gains 2 electrons what happens to the size of the red atom?

When atoms react to form \_\_\_\_\_ the total number of electrons that are \_\_\_\_\_ must always be equal to the total number of electrons that are \_\_\_\_\_.

Formula:

+ ion

- ion

Total:

Total:

Formula:

+ ion

- ion

Total:

Total:

Formula:

+ ion

- ion

Total:

Total:

# Ionic Formulas Chemthink Answers

**DP Hallahan**



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