

### Part C: Electron Configuration

12. How many electrons can each level hold? 1st = \_\_\_\_\_ 2nd = \_\_\_\_\_ 3rd = \_\_\_\_\_

13. What term is used for the electrons in the outermost shell or energy level? \_\_\_\_\_

14. Scientists use two types of diagrams to show the electron configuration for atoms. Follow your teacher's directions to complete the diagrams.

Sulfur

Atomic # = 16

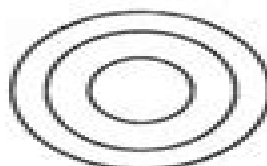
Atomic Mass = 32

Protons = \_\_\_\_\_

Neutrons = \_\_\_\_\_

Electron = \_\_\_\_\_

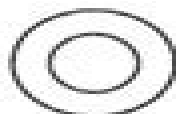
Bohr Diagram  
Shows all electrons



Lewis Structure  
Shows valence electrons

S

15. Calculate the missing information and then draw the Bohr Diagram and Lewis Structure for each element.



Li

Atomic # = 3

Mass # = 7

# of P = \_\_\_\_\_

# of N = \_\_\_\_\_

# of E = \_\_\_\_\_



Ne

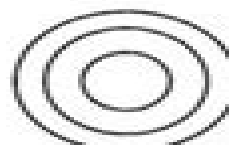
Atomic # = 10

Mass # = 20

# of P = \_\_\_\_\_

# of N = \_\_\_\_\_

# of E = \_\_\_\_\_



Mg

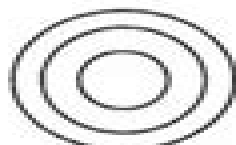
Atomic # = 12

Mass # = 24

# of P = \_\_\_\_\_

# of N = \_\_\_\_\_

# of E = \_\_\_\_\_



Cl

Atomic # = 17

Mass # = 35

# of P = \_\_\_\_\_

# of N = \_\_\_\_\_

# of E = \_\_\_\_\_



He

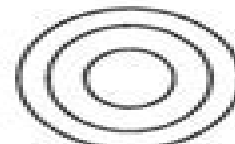
Atomic # = 2

Mass # = 4

# of P = \_\_\_\_\_

# of N = \_\_\_\_\_

# of E = \_\_\_\_\_



Si

Atomic # = 14

Mass # = 28

# of P = \_\_\_\_\_

# of N = \_\_\_\_\_

# of E = \_\_\_\_\_

16. Answer the questions below based on the elements in question #15.

(1) Which elements had a filled outermost shell? \_\_\_\_\_

(2) Which element would be most likely to lose electrons in a chemical bond? \_\_\_\_\_

(3) Which element would be most likely to gain electrons in a chemical bond? \_\_\_\_\_

(4) Which elements are not likely to bond with other elements? \_\_\_\_\_ Why? \_\_\_\_\_

# Ionic And Covalent Bonds Worksheet 7 2 Answers

**Chandan Sengupta**

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the author's name.

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## **Table of Contents Ionic And Covalent Bonds Worksheet 7 2 Answers**

1. Understanding the eBook Ionic And Covalent Bonds Worksheet 7 2 Answers
  - The Rise of Digital Reading Ionic And Covalent Bonds Worksheet 7 2 Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Ionic And Covalent Bonds Worksheet 7 2 Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Ionic And Covalent Bonds Worksheet 7 2 Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Ionic And Covalent Bonds Worksheet 7 2 Answers
  - Personalized Recommendations
  - Ionic And Covalent Bonds Worksheet 7 2 Answers User Reviews and Ratings
  - Ionic And Covalent Bonds Worksheet 7 2 Answers and Bestseller Lists
5. Accessing Ionic And Covalent Bonds Worksheet 7 2 Answers Free and Paid eBooks
  - Ionic And Covalent Bonds Worksheet 7 2 Answers Public Domain eBooks
  - Ionic And Covalent Bonds Worksheet 7 2 Answers eBook Subscription Services
  - Ionic And Covalent Bonds Worksheet 7 2 Answers Budget-Friendly Options

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  - ePub, PDF, MOBI, and More
  - Ionic And Covalent Bonds Worksheet 7 2 Answers Compatibility with Devices
  - Ionic And Covalent Bonds Worksheet 7 2 Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Ionic And Covalent Bonds Worksheet 7 2 Answers
  - Highlighting and Note-Taking Ionic And Covalent Bonds Worksheet 7 2 Answers
  - Interactive Elements Ionic And Covalent Bonds Worksheet 7 2 Answers
8. Staying Engaged with Ionic And Covalent Bonds Worksheet 7 2 Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Ionic And Covalent Bonds Worksheet 7 2 Answers
9. Balancing eBooks and Physical Books Ionic And Covalent Bonds Worksheet 7 2 Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Ionic And Covalent Bonds Worksheet 7 2 Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Ionic And Covalent Bonds Worksheet 7 2 Answers
  - Setting Reading Goals Ionic And Covalent Bonds Worksheet 7 2 Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ionic And Covalent Bonds Worksheet 7 2 Answers
  - Fact-Checking eBook Content of Ionic And Covalent Bonds Worksheet 7 2 Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

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