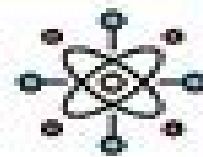


CHEMICAL BONDS



IONIC, COVALENT, & METALLIC

FORMS OF CHEMICAL BONDS
Bonding: combination of atoms

Atoms combine to form molecules or compounds by sharing or transferring electrons. This process is called bonding. The resulting molecule or compound has properties different from the individual atoms.

Types of Bonds:

- Covalent Bond:** Formed by sharing electrons between two atoms. Example: H_2 , H_2O , CO_2 .
- Ionic Bond:** Formed by the transfer of electrons from one atom to another, creating oppositely charged ions that attract each other. Example: $NaCl$, MgO .
- Metallic Bond:** Formed by a sea of delocalized electrons surrounding positive metal ions. Example: Fe , Cu , Al .

Covalent, Ionic, and Metallic Bonding

Covalent bonding occurs when two non-metal atoms share their outer electrons to form a stable molecule. Ionic bonding occurs when a metal atom transfers its outer electrons to a non-metal atom, forming oppositely charged ions. Metallic bonding occurs in metals, where the outer electrons are delocalized and form a 'sea of electrons' around the positive metal ions.

Science Real

Diagram of Ionic Bonding: Shows a red circle with a minus sign ($-$) and a yellow circle with a plus sign ($+$) connected by a dashed line, representing an ionic bond.

Diagram of Covalent Bonding: Shows two atoms, each with a nucleus and a shell of electrons. A red arrow points to the shared electron pair between the two atoms, representing a covalent bond.

Diagram of Metallic Bonding: Shows a cluster of positive metal ions (yellow circles with plus signs) surrounded by a sea of delocalized electrons (blue dots).

Text on the right:

When two atoms share a pair of electrons, they form a covalent bond. This process is called covalent bonding. The resulting molecule has properties different from the individual atoms.

When a metal atom transfers its outer electrons to a non-metal atom, they form oppositely charged ions. These ions are attracted to each other, forming an ionic bond. The resulting compound has properties different from the individual atoms.

In metals, the outer electrons are delocalized and form a 'sea of electrons' around the positive metal ions. This is called metallic bonding. The resulting metal has properties different from the individual atoms.

GUIDED READING ACTIVITY

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Goyal Brothers Prakashan



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California Edition Michael J. Padilla, 2007 **Chemical Interactions**, 2005 How to Study Science Fred Drewes, 2000 This text aims to help students get the most out of their science course by giving them suggestions on notetaking managing study time and taking tests A multidisciplinary approach is taken including examples from biology chemistry physics geology and meteorology Understanding Earth Student Study Guide Peter L. Kresan, Reed Mencke, 2006-05-03 The guide helps students prepare for lectures and exams with a heavy emphasis on utilizing the book's Web resources **Chemistry**

Questions - Answers for Intermediate Mr. Ajeet Singh, 2018-05-13 A question answer formatted book designed for intermediate level chemistry students covering theory based and numerical questions to aid learning and exam preparation

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Engineering ,1876 *Essentials of Biology* Holt Rinehart & Winston, 1998

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